

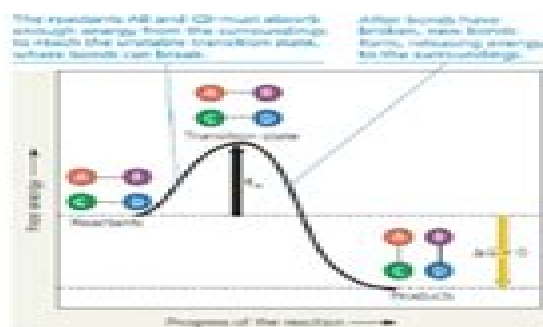
Study Guide: Enzymes

Enzymes are chemical agents that speed up a chemical reaction without being consumed. It usually ends with -ase.

A spontaneous chemical reaction occurs without any requirement for outside energy, but it may occur so slowly that it is imperceptible. For example, even though the hydrolysis of sucrose (table sugar) to glucose and fructose is exergonic, occurring spontaneously with a release of free energy ($G = -7$ kcal/mol), a solution of sucrose dissolved in sterile water will sit for years at room temperature with no appreciable hydrolysis. However, if we add a small amount of the enzyme sucrase to the solution, then all the sucrose may be hydrolyzed within seconds, as shown here:



How exactly do enzymes speed up the chemical reaction? Let's check the graph below.



- Chemical reaction involves breaking or forming a bond. e.g., sucrose $\gg$ glucose and fructose
- Changing one molecule into another generally involves contorting the starting molecule into a highly unstable state before the reaction can proceed.
- Contorting the reactants is like trying to insert a key into the key holder. The keyring is highly unstable in its opened form but returns to a stable state once the key is threaded all the way onto the ring.
- To reach the contorted, molecules need to absorb energy from their surroundings
- Energy required to contort the reactant molecules is called *Activation Energy*. It is supplied with thermal energy.
- The energy accelerates the reactant molecules and collides more often causing the breakage of a bond.



Study Guide On Enzymes

**Marcia L. Gillette, Ira Blei, Wendy
Gloffke, George Odian**



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