

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and State of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	8 molecules	1, 4 molecules	2, O_2	3, 4 molecules NO
1.00 mol	2.00 mol	4, 2.00 mol	5, none	6, none
4.00 mol	4.00 mol	7, 4.00 mol	8, NO	9, 2.00 mol O_2
1.00 mol	7.00 mol	10, 7.00 mol	11, NO	12, 1.50 mol O_2
1.00 mol	4.00 mol	13, 2.00 mol	14, O_2	15, 2.00 mol NO
0.500 mol	0.200 mol	16, 0.200 mol	17, NO	18, 0.400 mol O_2
32.00 g	60.02 g	19, 62.02 g	20, none	21, none
16.00 g	60.02 g	22, 46.01 g	23, O_2	24, 58.12 g NO
16.00 g	20.00 g	25, 28.76 g	26, O_2	27, 1.24 g NO

balanced equation mole ratio = 2 mol NO/1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO} / 30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $0.6664 \text{ mol } \text{NO} / 0.5125 \text{ mol } \text{O}_2 = 2.132 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of NO: O_2 is larger than the balanced equation mole ratio of NO: O_2 , there is an excess of NO; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

Excess NO = $20.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 7.24 \text{ g } \text{NO}$

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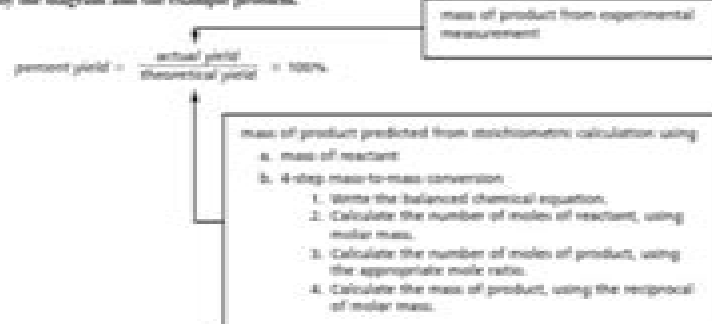
CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown $7.42 \text{ g } \text{Ga}_2\text{O}_3$
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 7.00 \text{ g } \text{Ga} \times 1 \text{ mol } \text{Ga} / 69.72 \text{ g } \text{Ga}$
- number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}} = 0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}}$
- theoretical yield: $0.100 \text{ mol } \text{Ga}_2\text{O}_3 \times \frac{187.44 \text{ g } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}_2\text{O}_3} = 18.74 \text{ g } \text{Ga}_2\text{O}_3$ correct
- percent yield: $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{18.74 \text{ g } \text{Ga}_2\text{O}_3} \times 100 = 7.42 \text{ g } \text{Ga}_2\text{O}_3 / 18.74 \text{ g } \text{Ga}_2\text{O}_3 \times 100$

Study Guide For Content Mastery Limiting Reactants

F Rizvi



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Table of Contents Study Guide For Content Mastery Limiting Reactants

1. Understanding the eBook Study Guide For Content Mastery Limiting Reactants
 - The Rise of Digital Reading Study Guide For Content Mastery Limiting Reactants
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Limiting Reactants
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Mastery Limiting Reactants
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Mastery Limiting Reactants
 - Personalized Recommendations
 - Study Guide For Content Mastery Limiting Reactants User Reviews and Ratings

- Study Guide For Content Mastery Limiting Reactants and Bestseller Lists
- 5. Accessing Study Guide For Content Mastery Limiting Reactants Free and Paid eBooks
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 - Study Guide For Content Mastery Limiting Reactants Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Study Guide For Content Mastery Limiting Reactants
 - Highlighting and Note-Taking Study Guide For Content Mastery Limiting Reactants
 - Interactive Elements Study Guide For Content Mastery Limiting Reactants
- 8. Staying Engaged with Study Guide For Content Mastery Limiting Reactants
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Mastery Limiting Reactants
- 9. Balancing eBooks and Physical Books Study Guide For Content Mastery Limiting Reactants
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Content Mastery Limiting Reactants
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Content Mastery Limiting Reactants
 - Setting Reading Goals Study Guide For Content Mastery Limiting Reactants
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Content Mastery Limiting Reactants
 - Fact-Checking eBook Content of Study Guide For Content Mastery Limiting Reactants
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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