# Assignment 1. Questions from chapters 1 and 2 of McMurry and Fay Question numbers are from the fourth edition.

- 1.1 (a) Cd (b) Sb (c) Am
- 1.2 (a) silver (b) rhodium (c) rhenium (d) cesium (e) argon (f) arsenic
- 1.3 (a) Ti, metal (b) Te, semimetal (c) Se, nonmetal (d) Sc, metal (e) At, semimetal (f) Ar, nonmetal
- 1.4 The three Acoinage metals@ are copper (Cu), silver (Ag), and gold (Au).
- 1.5 (a) The decimal point must be shifted ten places to the right so the exponent is \$10. The result is 3.72 x 10<sup>-10</sup> m.
  - (b) The decimal point must be shifted eleven places to the left so the exponent is 11. The result is 1.5 x 10<sup>11</sup> m.
- 1.6 (a) microgram(b) decimeter (c) picosecond
  - (d) kiloampere (e) millimole
- 1.7  $^{\circ}$ C =  $\frac{5}{9}$ x ( $^{\circ}$ F  $_{-}$ 32) =  $\frac{5}{9}$ x (98.6  $_{-}$ 32) = 37.0 $^{\circ}$ C K =  $^{\circ}$ C + 273.15 = 37.0 + 273.15 = 310.2 K
- 1.8 (a)  $K = {}^{\circ}C + 273.15 = 878 + 273.15 = 195.15 K = 195 K$ 
  - (b)  $^{\circ}F = (\frac{9}{5} \times ^{\circ}C) + 32 = (\frac{9}{5} \times 158) + 32 = 316.4^{\circ}F = 316^{\circ}F$
  - (c)  $^{\circ}$ C = K  $\approx 273.15 = 375 \approx 273.15 = 101.85 ^{\circ}$ C =  $102 ^{\circ}$ C  $^{\circ}$ F =  $(\frac{9}{5} \times ^{\circ}$ C) +  $32 = (\frac{9}{5} \times 101.85) + 32 = 215.33 ^{\circ}$ F =  $215 ^{\circ}$ F
- 1.9  $d = \frac{m}{V} = \frac{27.43 \text{ g}}{12.40 \text{ cm}^3} = 2.212 \text{ g/cm}^3$
- 1.10 volume = 9.37 g x  $\frac{1 \text{ mL}}{1.483 \text{ g}}$  = 6.32 mL
- 1.11 The actual mass of the bottle and the acetone = 38.0015 g + 0.7791 g = 38.7806 g. The measured values are 38.7798 g. 38.7795 g. and 38.7801 g. These values are both close to each other and close to the actual mass. Therefore the results are both precise and accurate.

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