Name : _____ Date : _____

Score : _____

Solubility Rules Practice Worksheet

1. Classify each substance as being soluble or insoluble in water.

I. $Mg(PO_4)_2$ – _____

II. KOH – _____

III. $NiCl_2$ – _____

IV. NH_4OH – _____

V. Hg_2SO_4 – _____

VI. KBr – _____

VII. $Pb(CO_3)$ - _____

VIII. PbI_2 - _____

IX. $BaSO_4$ - _____

X. $NiCl_2$ – _____

2. Show the ions that formed the following compounds:

I. $Zn_3(PO_4)_2$

II. Al_2S_3

III. Iron (III) sulfide

IV. Ammonium cyanide

3. Form 4 water-soluble compounds by combining ions from the ions below:

Cl^- CO_3^{2-} PO_4^{3-} Li^+ Sr^{2+}

4. Identify the precipitate in the following reaction. Circle the correct answer.

a. $Li_2CO_3 + Co(CH_3COO)_2 \rightarrow 2\ LiCH_3COO + CoCO_3$

b. $Pb(NO_3)_2 + Li_2SO_4 \rightarrow PbSO_4 + 2\ LiNO_3$

5. Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formula	Name	Solubility
$Zn_3(PO_4)_2$		
	Silver bromide	
KNO_3		
	Aluminum sulfide	
	Silver acetate	

Solubility Practice Problems And Answers

Madan R.L.

Solubility Practice Problems And Answers:

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profession in India tracing its development through education industrial growth and organizational progress The book provides a detailed introduction to various dosage forms offering clarity on their classification definitions and relevance in drug delivery Emphasis is placed on prescriptions their structure handling and potential errors helping students understand the importance of accuracy in pharmaceutical practice The section on posology explores dosage calculation principles particularly for pediatric patients incorporating methods based on age weight and body surface area Pharmaceutical calculations are addressed with practical examples involving percentage solutions proof spirit isotonic solutions and molecular weights ensuring mathematical accuracy in formulation Detailed attention is given to powders including classifications such as effervescent hygroscopic and eutectic mixtures along with methods like geometric dilution for uniform mixing Liquid dosage forms are elaborated with their pros and cons common excipients and techniques to enhance solubility A full chapter on monophasic liquids describes the preparation and use of various formulations such as syrups elixirs enemas and lotions The book also delves into suspensions covering flocculation deflocculation and methods to enhance stability Similarly the chapter on emulsions explores types emulsifying agents preparation methods and stability challenges Suppositories are discussed in terms of types preparation techniques displacement values and evaluation parameters The book also highlights different types of pharmaceutical incompatibilities physical chemical and therapeutic with practical examples A robust section on semisolid dosage forms covers the formulation and evaluation of ointments creams pastes and gels while addressing drug penetration mechanisms and formulation variables This textbook serves as a vital guide for students offering a strong theoretical foundation supported by practical insights ensuring they are well prepared for professional pharmacy practice *General Organic and Biological Chemistry* Kenneth W. Raymond,2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

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