

CHAPTER 6

STUDY GUIDE

Ionic Compounds and Metals

Lesson 6.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) Chemical Bond. Such an attachment may form by the attraction of the positively charged (2) Nucleus of one atom for the negatively charged (3) Electron of another atom, or by the attraction of charged atoms, which are called (4) Ions. The attractions may also involve (5) Valence electrons, which are the electrons in the outermost (6) Energy level. The (7) Noble Gas are a family of elements that have very little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a stable (8) Octet. The relatively stable electron structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9) Pseudo-Noble gas formations.

For each statement below, write *true* or *false*.

- | | |
|--------------|--|
| <u>False</u> | 10. A positively charged ion is called an anion. |
| <u>True</u> | 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge. |
| <u>False</u> | 12. Elements tend to react so that they acquire the electron structure of a halogen. |
| <u>True</u> | 13. A sodium atom tends to lose one electron when it reacts. |
| <u>True</u> | 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation. |
| <u>False</u> | 15. A Cl^- ion is an example of a cation. |
| <u>True</u> | 16. The ending <i>-ide</i> is used to designate an anion. |
| <u>False</u> | 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions. |

Study Guide Ionic Compounds Metals Answers

Martha Mackin



Study Guide Ionic Compounds Metals Answers:

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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