

Chapter 7: Ionic Compounds and Metals

7.1 Ion Formation p. 206

1. A chemical bond is the force that holds two atoms together.
2. Chemical bonds form by the attraction between the positive nucleus of one atom and the negative electrons of another atom.
3. Atom's try to form the octet—the stable arrangement of eight valence electrons in the outer energy level—by gaining or losing valence electrons.
4. A positively charged ion is called a cation.
5. An example is when sodium atom loses one valence electron to become a cation ion.
6. Metals are reactive because they lose valence electrons easily.
7. Transition metals commonly form 2+ or 3+ ions, but can form greater than 3+ ions.
8. Other relatively stable electron arrangements are referred to as pseudo-noble gas configurations.
9. An anion is a negatively charged ion.
10. An example would be chlorine gaining an electron to become a chlorine ion.
11. Nonmetal ions gain the number of electrons required to fill an octet.
12. Some nonmetals can gain or lose electrons to complete an octet.

7.2 Ionic Bonds and Ionic Compounds p.210

1. The electrostatic force that holds oppositely charged particles together in an ionic compound is called an ionic bond.
2. Compounds that contain ionic bonds are called ionic compounds.
3. Binary ionic compounds contain only two different elements—a metallic cation and a nonmetallic anion.
4. Formation of an ionic bond (table 7.4 p.211)
5. Positive and negative ions exist in a ratio determined by the number of electrons transferred from the metal atom to the non-metal atom.
6. The repeating pattern of particle packing in an ionic compound is called an ionic crystal.
7. The strong attractions among the positive and negative ions result in the formation of the crystal lattice.
8. A crystal lattice is the three-dimensional geometric arrangement of particles, and is responsible for the structure of many minerals.
9. Melting point, boiling point, and hardness depend on the strength of the attraction.
10. Melting and boiling points of some ionic compounds table 7.5 p.215.
11. In a solid, ions are locked into position and electrons cannot flow freely—solid ions are poor conductors of electricity.
12. Liquid ions or ions in aqueous solution have electrons that are free to move, so they conduct electricity easily.
13. An ion in aqueous solution that conducts electricity is an electrolyte.
14. Reactions that absorb energy are endothermic.
15. Reactions that release energy are exothermic.
16. The energy required to separate 1 mol of ions in an ionic compound is referred to as the lattice energy.
17. Lattice energy is directly related to the size of the ions that are bonded.

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1. Understanding the eBook Study Guide Metals Answers
 - The Rise of Digital Reading Study Guide Metals Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Metals Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Metals Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Metals Answers
 - Personalized Recommendations
 - Study Guide Metals Answers User Reviews and Ratings
 - Study Guide Metals Answers and Bestseller Lists
5. Accessing Study Guide Metals Answers Free and Paid eBooks
 - Study Guide Metals Answers Public Domain eBooks
 - Study Guide Metals Answers eBook Subscription Services

- Study Guide Metals Answers Budget-Friendly Options
- 6. Navigating Study Guide Metals Answers eBook Formats
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 - Study Guide Metals Answers Compatibility with Devices
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- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Metals Answers
 - Highlighting and Note-Taking Study Guide Metals Answers
 - Interactive Elements Study Guide Metals Answers
- 8. Staying Engaged with Study Guide Metals Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Metals Answers
- 9. Balancing eBooks and Physical Books Study Guide Metals Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Metals Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Metals Answers
 - Setting Reading Goals Study Guide Metals Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Metals Answers
 - Fact-Checking eBook Content of Study Guide Metals Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
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oggi ho pensato di cuocere questa pietanza in airy fryer con pochissimo olio il risultato è continua a leggere ricette di secondi
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