

(6.3-4)

$T = 293 \text{ K}$ $\text{CH}_3\text{OH}(A)$ $\text{H}_2\text{O}(B)$ $M_B = 18.02$

Appendix A.2 $\mu_B = 1.005 \times 10^{-3} \text{ kg/m}\cdot\text{s}$

Table 6.3-2

$$V_A = 0.0148 + 4(0.0037) + 1(0.0074) = 0.0370 \text{ m}^3/\text{kg mol}$$

Eq. (6.3-9)

$$D_{AB} = \frac{1.173 \times 10^{-16} (p_B M_B)^{1/2} T}{\mu_B V_A^{0.6}}$$

$$p_B = 2.6$$

(a)

$$= \frac{1.173 \times 10^{-16} (2.6 \times 18.02)^{1/2} (293)}{1.005 \times 10^{-3} (0.0370)^{0.6}}$$

$$\boxed{D_{AB} = 1.692 \times 10^{-9} \text{ m}^2/\text{s}}$$

(b) $T = 288 \text{ K}$ A.2 $\mu_B = 1.1404 \times 10^{-3} \text{ kg/m}\cdot\text{s}$

$$D_{AB} \propto \frac{T}{\mu_B}$$

$$D_{AB \text{ eff}} = 1.26 \times 10^{-9} \text{ m}^2/\text{s} \quad (288 \text{ K})$$

$$D_{AB} = 1.26 \times 10^{-9} \left(\frac{293}{288} \right) \left(\frac{1.1404}{1.005} \right)$$

$$\boxed{D_{AB} = 1.455 \times 10^{-9} \text{ m}^2/\text{s}}$$

(6.4-1)

Table 6.4-1 $M_A = 482,700$

$T = 298 \text{ K}$

$$\boxed{D_{AB \text{ eff}} = 4.01 \times 10^{-11} \text{ m}^2/\text{s}}$$

A.2 $\mu = 0.8937 \times 10^{-3} \text{ kg/m}\cdot\text{s} \quad (298 \text{ K})$

Eq. (6.4-1)

$$D_{AB} = \frac{9.40 \times 10^{-15} T}{\mu (M_A)^{1/3}}$$

$$= \frac{9.40 \times 10^{-15} (298)}{0.8937 \times 10^{-3} (482,700)^{1/3}}$$

$$\boxed{D_{AB} = 3.995 \times 10^{-11} \text{ m}^2/\text{s}}$$

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Table of Contents Solution Manual Geankoplis

1. Understanding the eBook Solution Manual Geankoplis
 - The Rise of Digital Reading Solution Manual Geankoplis
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual Geankoplis
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual Geankoplis
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Manual Geankoplis
 - Personalized Recommendations
 - Solution Manual Geankoplis User Reviews and Ratings
 - Solution Manual Geankoplis and Bestseller Lists
5. Accessing Solution Manual Geankoplis Free and Paid eBooks
 - Solution Manual Geankoplis Public Domain eBooks

- Solution Manual Geankoplis eBook Subscription Services
- Solution Manual Geankoplis Budget-Friendly Options
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 - Solution Manual Geankoplis Compatibility with Devices
 - Solution Manual Geankoplis Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Manual Geankoplis
 - Highlighting and Note-Taking Solution Manual Geankoplis
 - Interactive Elements Solution Manual Geankoplis
- 8. Staying Engaged with Solution Manual Geankoplis
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual Geankoplis
- 9. Balancing eBooks and Physical Books Solution Manual Geankoplis
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual Geankoplis
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solution Manual Geankoplis
 - Setting Reading Goals Solution Manual Geankoplis
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solution Manual Geankoplis
 - Fact-Checking eBook Content of Solution Manual Geankoplis
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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