# LAB 13: SOLUBILITY PRODUCT CONSTANT



#### Goals

- 1. Learn of solubility product constants
- 2. Estimate a solubility constant
- Observe that a solubility constant is a special case of an equilibrium constant

# **Materials and Equipment**

Reagents:

0.1 M calcium nitrate,  $Co(NO_s)$ ,

0.1 M sodium oxalate,  $Na_sC_sO_a$ Pipet, mini
Toothpick, plastic
Reaction plate, 96 well

# **Materials Not Included**

A sheet of black paper Distilled water

#### Introduction

dissociate to form ions when they dissolve in water. An example is sodium chloride:

(Equation 13.1) 
$$NaCl(s) \rightarrow Na^{*}(aq) + Cl^{*}(aq)$$

Solid NaCl dissociates to become National Cl. ions. The National cations are attracted to the negative "ends" of water molecules, and the Cl. anions are attracted to the positive "ends" of water molecules.

In other compounds the attraction between the cations and the anions is stronger, and there is little dissociation. These are sparingly or slightly soluble salts. An example is silver chloride, AgCI. A beaker of this solution often has solid AgCI crystals in the bottom. The crystals are

in equilibrium with the Ag' and CI ions that are in the solution:

(Equation 13.2) 
$$AgCl(s) \leftrightarrow Ag'(aq) + Cl(aq)$$

In this reaction, the rate of AgCI dissolving is equal to the rate of Ag and CI ions combining to form solid AgCI. In other words, the reaction is going both ways at the same time and at the same rate. It is said to be in equilibrium. The equilibrium constant is:

(Equation 13.3) 
$$K = [Ag^{-}]/CT$$

AgCI(s) is not considered in the equilibrium constant because it is a solid.

We call this special equilibrium constant the solubility product constant, K...

Some of the beautiful structures you may have seen in caves are composed of sparingly soluble calcium carbonate. This material slightly dissolves in water, and later precipitates when the water in the solution evaporates or the solution cools.

Another sparingly soluble salt is calcium oxalate, which sometimes precipitates in our kidneys as kidney stones. These can be very painful and often require medical treatment. We will study calcium oxalate,  $CaC_{i}O_{j}$ , in this experiment. The solubility product constant for this is:

(Equation 13.4) 
$$K_{ap} = [Ca^2][C_2O_2^2]$$

# **Solubility Lab Answer Key**

United States. National Bureau of Standards

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