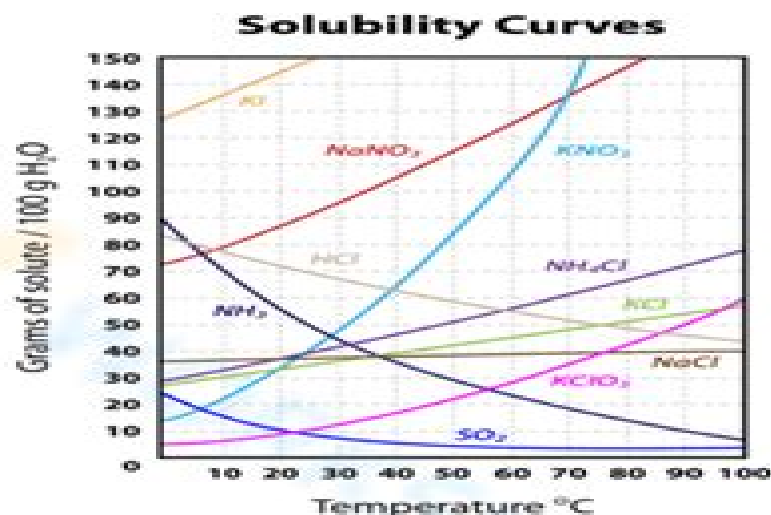


WS 6.1: Interpreting Solubility Curves

Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 mL.



- What is the solvent in this Solubility Curve?
water
- Saturation points are given per how much solvent?
100 ml or 100 g
- What is the saturation point of NaNO₃ at 10°C, 40°C, & 80°C?
@ 10°C = ~79 g
@ 40°C = ~105 g
@ 80°C = ~146 g
- How much NH₃ can you dissolve in water at 10°C, 30°C, & 90°C?
@ 10°C = ~70 g
@ 30°C = ~45 g
@ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C, 50°C, & 90°C?
@ 20°C = ~33 g
@ 50°C = ~42 g
@ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C?
About 40 grams
- Except for KI What substance is the most soluble in water at 20°C?
NaNO₃
- Except for KI, what substance is the most soluble in water at 72°C?
KNO₃
- Look up The boiling points for SO₂, HCl and NH₃ and write them below:
B.P. SO₂ = -10°C
B.P. HCl = -85°C
B.P. NH₃ = -33°C
- Based on the boiling points you looked up, are SO₂, HCl & NH₃ solids, liquids, or gases at 0°C? (and anywhere else on the solubility curve)
gases

Solubility Curves Practice Problemsworksheet

Heather Hattori, Richard H. Langley

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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Table of Contents Solubility Curves Practice Problemsworksheet

1. Understanding the eBook Solubility Curves Practice Problemsworksheet
 - The Rise of Digital Reading Solubility Curves Practice Problemsworksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Curves Practice Problemsworksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Curves Practice Problemsworksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Curves Practice Problemsworksheet
 - Personalized Recommendations
 - Solubility Curves Practice Problemsworksheet User Reviews and Ratings
 - Solubility Curves Practice Problemsworksheet and Bestseller Lists
5. Accessing Solubility Curves Practice Problemsworksheet Free and Paid eBooks
 - Solubility Curves Practice Problemsworksheet Public Domain eBooks
 - Solubility Curves Practice Problemsworksheet eBook Subscription Services

- Solubility Curves Practice Problemsworksheet Budget-Friendly Options
- 6. Navigating Solubility Curves Practice Problemsworksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Solubility Curves Practice Problemsworksheet Compatibility with Devices
 - Solubility Curves Practice Problemsworksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solubility Curves Practice Problemsworksheet
 - Highlighting and Note-Taking Solubility Curves Practice Problemsworksheet
 - Interactive Elements Solubility Curves Practice Problemsworksheet
- 8. Staying Engaged with Solubility Curves Practice Problemsworksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Curves Practice Problemsworksheet
- 9. Balancing eBooks and Physical Books Solubility Curves Practice Problemsworksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Curves Practice Problemsworksheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solubility Curves Practice Problemsworksheet
 - Setting Reading Goals Solubility Curves Practice Problemsworksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solubility Curves Practice Problemsworksheet
 - Fact-Checking eBook Content of Solubility Curves Practice Problemsworksheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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