

1. (5 points) Table 4.8 (Snoeyink and Jenkins) shows the concentration of carbonate species for solutions with distilled water and 10^{-5}M solutions of H_2CO_3^* ; sodium bicarbonate, NaHCO_3 ; and sodium carbonate, Na_2CO_3 , in a system open to the atmosphere and in a closed system.
- Why does the concentration of H_2CO_3^* remain constant in the open system (across the different solutions) and decreases in a closed system?
 - For both systems, the concentration of H^+ decreases when comparing a solution of distilled water and 10^{-5}M solutions of H_2CO_3^* ; NaHCO_3 ; Na_2CO_3 with lowest concentration observed for Na_2CO_3 . The decrease in H^+ concentration is greater for a closed system than for an open system. Explain this difference.

TABLE 4-8 Comparison of Composition and pH of Open and Closed Carbonate Systems

Solution	Closed System (Fig. 4-17)						
	$[\text{H}^+]$	$[\text{OH}^-]$	$[\text{H}_2\text{CO}_3^*]$	$[\text{HCO}_3^-]$	$[\text{CO}_3^{2-}]$	pH	$C_{\text{T,CO}_3}$
$10^{-5}\text{ M H}_2\text{CO}_3^*$	$10^{-6.3}$	$10^{-7.7}$	$10^{-5.1}$	$10^{-6.7}$	$10^{-10.2}$	5.7	$10^{-5.1}$
10^{-5} M NaHCO_3	$10^{-7.8}$	$10^{-6.2}$	$10^{-6.4}$	10^{-5}	$10^{-10.4}$	7.6	$10^{-5.1}$
$10^{-5}\text{ M Na}_2\text{CO}_3$	10^{-9}	10^{-5}	$10^{-7.6}$	10^{-5}	$10^{-4.2}$	9.0	$10^{-5.1}$

Solution	Open System						
	$[\text{H}^+]$	$[\text{OH}^-]$	$[\text{H}_2\text{CO}_3^*]$	$[\text{HCO}_3^-]$	$[\text{CO}_3^{2-}]$	pH	$C_{\text{T,CO}_3}$
$10^{-5}\text{ M H}_2\text{CO}_3^*$	$10^{-6.3}$	$10^{-7.7}$	10^{-5}	$10^{-6.7}$	$10^{-10.2}$	5.7	$10^{-5.1}$
10^{-5} M NaHCO_3	$10^{-6.3}$	$10^{-7.7}$	10^{-5}	10^{-5}	$10^{-9.3}$	6.3	$10^{-5.1}$
$10^{-5}\text{ M Na}_2\text{CO}_3$	$10^{-6.3}$	$10^{-7.7}$	10^{-5}	$10^{-5.7}$	$10^{-8.3}$	6.6	$10^{-5.1}$

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