

Name: KEY

Valence Electrons Practice

The **valence electrons** are the electrons in the outermost energy level. They are always the outermost electrons.

Atoms that LOSE electrons are called **cations** and have an overall **positive** charge.

Atoms that GAIN electrons are called **anions** and have an overall **negative** charge.



Determine the element's number of valence electrons (# of electrons in the "outermost" shell).

Example: Carbon has 4 valence electrons, carbon 4

1. fluorine 7
2. lithium 1
3. phosphorus 5
4. francium 1
5. calcium 2
6. carbon 4
7. nitrogen 5

8. iodine 7
9. silicon 4
10. oxygen 6
11. argon 8
12. barium 2
13. potassium 1
14. aluminum 3

15. helium 2
16. hydrogen 1
17. magnesium 2
18. xenon 8
19. sulfur 6
20. boron 3
21. bromine 7

↑ 2,3

↓

22. What are valence electrons of an atom used for?
not part of Unit 2 exam

Determining how elements bond together

23. Which groups on the periodic table want to lose electrons? What kind(s) of elements are these groups?
Group 1, 2 elements, Alkali metals and Alkali earth metals

24. Which groups on the periodic table want to gain electrons? What kind(s) of elements are these groups?
Groups 5, 6, 7, non-metals

25. Why do the elements from question 23 want to lose electrons? Why do the elements from question 24 want to gain electrons? (HINT: The answer is the same for both questions.)

By losing 1 or 2 electrons these groups form more stable ion having the electron configuration of noble gases

26. How many valence electrons do the Alkaline Earth metals have? Are these elements very reactive? Explain your answer.
2, yes

27. What is the name of the atoms that lose electrons? What is their overall charge?

Cation (+) charge

28. What is the name of the atoms that gain electrons? What is their overall charge?

Anions (-) charge

Valence Electrons Assessment Key

Mr. Rohit Manglik

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the name bar.

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Table of Contents Valence Electrons Assessment Key

1. Understanding the eBook Valence Electrons Assessment Key
 - The Rise of Digital Reading Valence Electrons Assessment Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Valence Electrons Assessment Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Valence Electrons Assessment Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Valence Electrons Assessment Key
 - Personalized Recommendations

- Valence Electrons Assessment Key User Reviews and Ratings
- Valence Electrons Assessment Key and Bestseller Lists
- 5. Accessing Valence Electrons Assessment Key Free and Paid eBooks
 - Valence Electrons Assessment Key Public Domain eBooks
 - Valence Electrons Assessment Key eBook Subscription Services
 - Valence Electrons Assessment Key Budget-Friendly Options
- 6. Navigating Valence Electrons Assessment Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Valence Electrons Assessment Key Compatibility with Devices
 - Valence Electrons Assessment Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Valence Electrons Assessment Key
 - Highlighting and Note-Taking Valence Electrons Assessment Key
 - Interactive Elements Valence Electrons Assessment Key
- 8. Staying Engaged with Valence Electrons Assessment Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Valence Electrons Assessment Key
- 9. Balancing eBooks and Physical Books Valence Electrons Assessment Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Valence Electrons Assessment Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Valence Electrons Assessment Key
 - Setting Reading Goals Valence Electrons Assessment Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Valence Electrons Assessment Key
 - Fact-Checking eBook Content of Valence Electrons Assessment Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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