

DATE: 10/20/21
 SECTION/GROUP: #10
 NAME: Bailey Spell
 NAME: Gavin Sprinkle
 NAME: Ben Steadley

Investigation 3 Written Argument

Guiding Question: What is the concentration of acetic acid in vinegar?

Claim: My group and I claim that the concentration of acetic acid in vinegar is 0.83 M, 4.99% m/v

Evidence: Balanced equation: $1\text{CH}_3\text{COOH}(\text{aq}) + 1\text{NaOH}(\text{aq}) \gg> 1\text{NaCH}_3\text{COO}(\text{aq}) + 1\text{H}_2\text{O}(\text{l})$

Data Table

Data:	Trial 1:	Trial 2:	Trial 3:	Trial 4:
Molarity of NaOH used	0.400 M	0.400 M	0.400 M	0.400 M
Volume of Acetic Acid used	10.00 mL	10.00 mL	10.00 mL	10.00 mL
Initial Burette Reading (NaOH)	0.35 mL	21.15 mL	0.49 mL	21.21 mL
Final Burette Reading (NaOH)	21.15 mL	41.87 mL	21.21 mL	42.15 mL
Volume of NaOH used	20.80 mL	20.72 mL	20.72 mL	20.94 mL
Shade of Pink	Light pink	Light pink	Dark pink	Dark pink

Calculations of volume of NaOH used:

Trial 1: Initial Reading - Final Reading

21.15 mL NaOH - 0.35 mL of NaOH = 20.80 mL NaOH

Trial 2: Initial Reading - Final Reading

41.87 mL NaOH - 21.15 mL of NaOH = 20.72 mL NaOH

Trial 3: Initial Reading - Final Reading

21.21 mL NaOH - 0.49 mL NaOH = 20.72 mL NaOH

Trial 4: Initial Reading - Final Reading

42.15 mL NaOH - 21.21 mL NaOH = 20.94 mL NaOH

Calculations of moles of NaOH used:

Moles of NaOH used = Molarity x Volume in L

Make sure to change mL of NaOH into L of NaOH

Trial 1: $20.80 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02080 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02080 \text{ L NaOH} = 0.00832 \text{ mols NaOH}$

Trial 2: $20.72 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02072 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02072 \text{ L NaOH} = 0.00829 \text{ mols NaOH}$

Trial 3: $20.72 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02072 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02072 \text{ L NaOH} = 0.00829 \text{ mols NaOH}$

Trial 4: $20.94 \text{ mL NaOH} \times \frac{1 \text{ L NaOH}}{1000 \text{ mL NaOH}} = 0.02094 \text{ L NaOH} \gg> 0.400 \text{ M NaOH} \times 0.02094 \text{ L NaOH} = 0.00838 \text{ mols NaOH}$

Calculations of moles of Acetic Acid:

mols of Acetic Acid is same as mols of NaOH due to balanced equation

Make sure to change mL of volume of acetic acid used to L of acetic acid used $\gg> \gg$

$(10.00 \text{ mL} \times 1 \text{ L}) / 1000 \text{ mL} = 0.010 \text{ L}$

Molarity = mols/L

Trial 1: $0.00832 \text{ mols A.A.} / 0.010 \text{ L A.A.} = 0.83 \text{ M A.A.}$

Trial 2: $0.00829 \text{ mols A.A.} / 0.010 \text{ L A.A.} = 0.83 \text{ M A.A.}$

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