

SOLUTION & COLLIGATIVE PROPERTIES

OSMOTIC PRESSURE :

- (i) $\pi = \rho gh$ Where, ρ = density of soln., h = equilibrium height.
 (ii) **Vont – Hoff Formula** (For calculation of O.P.)
 $\pi = CST$

$$\pi = CRT = \frac{n}{V} RT \text{ (just like ideal gas equation)}$$

$$\begin{aligned} \therefore C &= \text{total conc. of all types of particles.} \\ &= C_1 + C_2 + C_3 + \dots \\ &= \frac{(n_1 + n_2 + n_3 + \dots)}{V} \end{aligned}$$

Note : If V_1 mL of C_1 conc. + V_2 mL of C_2 conc. are mixed.

$$\pi = \left(\frac{C_1 V_1 + C_2 V_2}{V_1 + V_2} \right) RT \quad ; \quad \pi = \left(\frac{\pi_1 V_1 + \pi_2 V_2}{RT} \right)$$

Type of solutions :

(a) **Isotonic solution** – Two solutions having same O.P.

$$\pi_1 = \pi_2 \text{ (at same temp.)}$$

(b) **Hyper tonic**– If $\pi_1 > \pi_2 \Rightarrow$ 1st solution is hypertonic solution w.r.t. 2nd solution.

(c) **Hypotonic** – 2nd solution is hypotonic w.r.t. 1st solution.

Abnormal Colligative Properties : (In case of association or dissociation)

VANT HOFF CORRECTION FACTOR (i) :

$$\begin{aligned} i &= \frac{\text{exp/observed / actual / abnormal value of colligative property}}{\text{Theoritical value of colligative property}} \\ &= \frac{\text{exp./observed no. of particles / conc.}}{\text{Theoritical no. of particles}} = \frac{\text{observed molality}}{\text{Theoritical molality}} \\ &= \frac{\text{theoretical molar mass(formula mass)}}{\text{experimental / observed molar mass(apparent molar mass)}} \end{aligned}$$

- $i > 1 \Rightarrow$ dissociation.
 ○ $i < 1 \Rightarrow$ association.

○ $i = \frac{\pi_{\text{exp.}}}{\pi_{\text{theor}}}$

$$\begin{aligned} \therefore \pi &= iCRT \\ \pi &= (i_1 C_1 + i_2 C_2 + i_3 C_3, \dots) RT \end{aligned}$$

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Table of Contents Solution And Colligative Properties Hsc

1. Understanding the eBook Solution And Colligative Properties Hsc
 - The Rise of Digital Reading Solution And Colligative Properties Hsc
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution And Colligative Properties Hsc
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution And Colligative Properties Hsc
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution And Colligative Properties Hsc
 - Personalized Recommendations
 - Solution And Colligative Properties Hsc User Reviews and Ratings
 - Solution And Colligative Properties Hsc and Bestseller Lists
5. Accessing Solution And Colligative Properties Hsc Free and Paid eBooks

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- Solution And Colligative Properties Hsc eBook Subscription Services
- Solution And Colligative Properties Hsc Budget-Friendly Options
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 - Solution And Colligative Properties Hsc Compatibility with Devices
 - Solution And Colligative Properties Hsc Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution And Colligative Properties Hsc
 - Highlighting and Note-Taking Solution And Colligative Properties Hsc
 - Interactive Elements Solution And Colligative Properties Hsc
- 8. Staying Engaged with Solution And Colligative Properties Hsc
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution And Colligative Properties Hsc
- 9. Balancing eBooks and Physical Books Solution And Colligative Properties Hsc
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution And Colligative Properties Hsc
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solution And Colligative Properties Hsc
 - Setting Reading Goals Solution And Colligative Properties Hsc
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solution And Colligative Properties Hsc
 - Fact-Checking eBook Content of Solution And Colligative Properties Hsc
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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