

Solubility Product, K_{sp} - Worksheet 1

1) Write the dissociation equation and the solubility product (K_{sp}) expression for the dissociation of each of the following substances.

- a) PbSO_4 (s)
b) $\text{Mg}(\text{OH})_2$ (s)

- c) $\text{Pb}(\text{NO}_3)_2$ (s)
d) AlCl_3 (s)

- e) Na_2S (s)
f) $\text{Ca}_3(\text{PO}_4)_2$ (s)

2) Write the solubility product expression for the dissociation of each of the following solids in water.

- a) sodium chloride
b) calcium carbonate

- c) silver sulfide
d) aluminum hydroxide

- e) magnesium sulfite
f) barium phosphate

3) The solubility of NiS is 1.0×10^{-12} mol/L. Calculate the value of K_{sp} for this compound.

4) The solubility of calcium hydroxide ($\text{Ca}(\text{OH})_2$) is 0.02 mol/L. Calculate the solubility product constant for this substance.

5) The solubility of silver sulfide (Ag_2S) is 1.3×10^{-6} mol/L. Calculate the K_{sp} value.

6) Experiments show that 0.0059 g of SrCO_3 will dissolve in 1.0 L of water. What is the K_{sp} for strontium carbonate?

7) The solubility of cobalt (II) sulfide is 3.8×10^{-3} g/L. Calculate the value of K_{sp} .

Solubility Product, K_{sp} - Worksheet 2

1. Calculate the solubility (in mol/L) for each of the following compounds:

- | | |
|-----------------------------|--------------------------------|
| a. FeS | $K_{sp} = 3.7 \times 10^{-19}$ |
| b. $\text{Mg}(\text{OH})_2$ | $K_{sp} = 9.0 \times 10^{-12}$ |
| c. Ag_2S | $K_{sp} = 1.6 \times 10^{-49}$ |
| d. AgCN | $K_{sp} = 2.0 \times 10^{-12}$ |

2. What is the solubility of silver iodide in grams per litre? ($K_{sp} = 8.3 \times 10^{-17}$)

3. The K_{sp} value for BaSO_4 is 1.5×10^{-9} . Calculate the number of moles of this salt that will dissolve in 1000 mL of water.

4. Calculate the mass of CuS that will dissolve in 1.0 litre of water. $K_{sp} = 4.0 \times 10^{-38}$

5. Calculate the mass of barium carbonate that will dissolve in 100 mL of water. The K_{sp} of BaCO_3 is 8.1×10^{-9} .

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