

(6.3-4)

$T = 293 \text{ K}$ $\text{CH}_3\text{OH}(A)$ $\text{H}_2\text{O}(B)$ $M_B = 18.02$

Appendix A.2 $\mu_B = 1.005 \times 10^{-3} \text{ kg/m}\cdot\text{s}$

Table 6.3-2

$$V_A = 0.0148 + 4(0.0037) + 1(0.0074) = 0.0370 \text{ m}^3/\text{kg mol}$$

Eq. (6.3-9)

$$D_{AB} = \frac{1.173 \times 10^{-16} (p_B M_B)^{1/2} T}{\mu_B V_A^{0.6}}$$

$$p_B = 2.6$$

(a)

$$= \frac{1.173 \times 10^{-16} (2.6 \times 18.02)^{1/2} (293)}{1.005 \times 10^{-3} (0.0370)^{0.6}}$$

$$\boxed{D_{AB} = 1.692 \times 10^{-9} \text{ m}^2/\text{s}}$$

(b) $T = 288 \text{ K}$ A.2 $\mu_B = 1.1404 \times 10^{-3} \text{ kg/m}\cdot\text{s}$

$$D_{AB} \propto \frac{T}{\mu_B}$$

$$D_{AB \text{ eff}} = 1.26 \times 10^{-9} \text{ m}^2/\text{s} \quad (288 \text{ K})$$

$$D_{AB} = 1.26 \times 10^{-9} \left(\frac{293}{288} \right) \left(\frac{1.1404}{1.005} \right)$$

$$\boxed{D_{AB} = 1.455 \times 10^{-9} \text{ m}^2/\text{s}}$$

(6.4-1)

Table 6.4-1 $M_A = 482,700$

$T = 298 \text{ K}$

$$\boxed{D_{AB \text{ eff}} = 4.01 \times 10^{-11} \text{ m}^2/\text{s}}$$

A.2 $\mu = 0.8937 \times 10^{-3} \text{ kg/m}\cdot\text{s} \quad (298 \text{ K})$

Eq. (6.4-1)

$$D_{AB} = \frac{9.40 \times 10^{-15} T}{\mu (M_A)^{1/3}}$$

$$= \frac{9.40 \times 10^{-15} (298)}{0.8937 \times 10^{-3} (482,700)^{1/3}}$$

$$\boxed{D_{AB} = 3.995 \times 10^{-11} \text{ m}^2/\text{s}}$$

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Michael R. Lindeburg



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