



# Chemical Changes

Read the passage below and answer the following questions:



1. What are the products of this chemical reaction?

Mercury and oxygen

2. What is the reactant? Mercury oxide

3. What is the physical state of the reactant and the products?

The reactant is a solid, the products are a liquid and a gas.

4. According to the law of conservation of mass, if 100g of mercury oxide is heated, how much mercury and oxygen will be produced?

100g



5. What products will be produced in this reaction? Mercury and oxygen

6. What mass of mercury will be produced? 80g

7. What are the physical states of the products in this reaction?

Mercury is a liquid, oxygen is a gas.

8. What is the physical state of the product in this reaction?

100g

9. What is the mass of the products in this reaction?

Mercury and oxygen

10. How many grams of product will be produced?

100g

Read the passage below and answer the following questions:

11. How many grams of  $\text{H}_2\text{O}$  will be produced if 10g of  $\text{H}_2$  reacts with 10g of  $\text{O}_2$ ?

10g of  $\text{H}_2\text{O}$  will be produced.

# Skills Directed A Section Chemical Changes Answers

**University of California (System)**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the University of California (System) text.

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