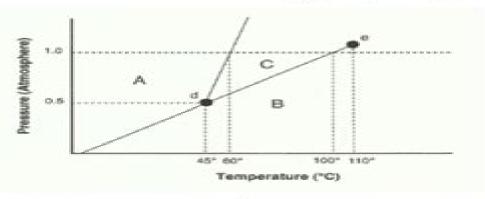


#### PHASE DIAGRAM WORKSHEET

#### Part A - Generic Phase Diagram.

Answer the questions below in relation to the following generic phase diagram.



- Which section represents the solid phase? A
- What section represents the liquid phase?
- What section represents the gas phase? B
- 4. What letter represents the triple point? d
  In your own words, what is the definition of a triple point?

It is the point where all three phases coexist at a specific temperature and pressure

- 5. What is this substance's normal melting point, at 1 atmosphere of pressure? 60°C
- What is this substance's normal boiling point, at 1 atmosphere of pressure? 100°C
- Above what temperature is it impossible to liquefy this substance, no matter what the pressure?
   110°C (This is known as the critical temperature)
- At what temperature and pressure do all three phases coexist? 45°C and 0.5 atm
- At a constant temperature, what would you do to cause this substance to change from the liquid phase to the solid phase? You would need to increase the pressure.
- 10 What does sublimation mean? The process of going from a solid to a gas directly, skipping the liquid phase altogether.

# **Skills Practice 4 Phase Diagrams Answers**

**ML Yell** 

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