SINGLE REPLACEMENT REACTIONS WORKSHEET

1.	Using the activity series table answer the following questions:
	(a) list two metals that will react with Fe ²⁺ ions
	(b) list two metals that will lose electrons to Pb2+
	(c) list two ions that will take electrons from Zn
	(d) list two metals that will react with HCI
	(e) identify the best oxidizing agent (i.e. electron grabber)
2.	Using the activity series, how do you decide if a reaction will or will not take place between two species?
3.	Describe what you would see when a piece of Mg metal is placed in an aqueous solution of copper(II) sulfate. Writ the net ionic equation for the reaction that takes placed.
4.	What would happen to a silver dollar when placed in a 3 M HCI solution? Justify your answer.
5.	Given an ionic equation, how can you identify the spectator ions of a reaction?
6.	Do aluminum pop cans make a suitable containers for hydrochloric acid solutions? Write an equation that supports your answer.
7.	
8.	Rank order the following species in order of their increasing ability to lose electrons: Ni , Cu*, Cl, K, Mg
9.	What happens when a piece of calcium metal is placed in water? Write the molecular equation that supports your prediction.

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Table of Contents Skill Practice 28 Answers On Single Replacement

- 1. Understanding the eBook Skill Practice 28 Answers On Single Replacement
 - The Rise of Digital Reading Skill Practice 28 Answers On Single Replacement
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Skill Practice 28 Answers On Single Replacement
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 28 Answers On Single Replacement
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Skill Practice 28 Answers On Single Replacement

- Personalized Recommendations
- Skill Practice 28 Answers On Single Replacement User Reviews and Ratings
- Skill Practice 28 Answers On Single Replacement and Bestseller Lists
- 5. Accessing Skill Practice 28 Answers On Single Replacement Free and Paid eBooks
 - Skill Practice 28 Answers On Single Replacement Public Domain eBooks
 - Skill Practice 28 Answers On Single Replacement eBook Subscription Services
 - Skill Practice 28 Answers On Single Replacement Budget-Friendly Options
- 6. Navigating Skill Practice 28 Answers On Single Replacement eBook Formats
 - o ePub, PDF, MOBI, and More
 - Skill Practice 28 Answers On Single Replacement Compatibility with Devices
 - Skill Practice 28 Answers On Single Replacement Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 28 Answers On Single Replacement
 - Highlighting and Note-Taking Skill Practice 28 Answers On Single Replacement
 - Interactive Elements Skill Practice 28 Answers On Single Replacement
- 8. Staying Engaged with Skill Practice 28 Answers On Single Replacement
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 28 Answers On Single Replacement
- 9. Balancing eBooks and Physical Books Skill Practice 28 Answers On Single Replacement
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 28 Answers On Single Replacement
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 28 Answers On Single Replacement
 - Setting Reading Goals Skill Practice 28 Answers On Single Replacement
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 28 Answers On Single Replacement

- Fact-Checking eBook Content of Skill Practice 28 Answers On Single Replacement
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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