

Moles and Chemical Formulas

11

A. Finding the Simplest Formula

1. Mass of empty crucible + cover
2. Initial appearance of the magnesium
3. Mass of crucible + cover + magnesium
4. Mass of crucible + cover + oxide product

22.842	g
23.483	
23.073	g
29.317	g

Calculations

5. Mass of magnesium
6. Mass of magnesium compound
7. Mass of oxygen in the product
8. Moles of Mg

0.2305	g
6.445	g
6.245	g
.0095	g

(Show calculations.)

$$2305 \text{ g/mg} \times \frac{1 \text{ mole mg}}{24.31 \text{ g/mg}} = .009481$$

9. Moles of O

(Show calculations.)

$$6.245 \text{ g} \times \frac{1 \text{ mole}}{16.00} = .3903$$

10. Which number of moles (Mg or O) is smaller Mole of Mg
moles of Mg (rounded)

Skill Practice 31 Moles And Formulas Answers

AW Rasmussen



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The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math. The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving. The book's unique focus on actual chemical practice, extensive study tools, and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors. This special PowerPoint tour of the text was created by Don Wink. http://www.bfwpub.com/pdfs/wink/POC/PowerPoint_Final.ppt 832KB The Software Finder, 1983 **The Software Encyclopedia**, 1986 *Holt Chemistry* Ralph Thomas Myers, 2004 **R.R. Bowker's Software for Schools**, 1987 Chemistry in the Community American Chemical Society, 2006-01-31 This laboratory-based text centres itself around decision-making activities where students apply their chemistry knowledge to realistic situations. This fifth edition includes more photographs, new drawings, and new design. **Chemical Principles** Peter Atkins, Loretta Jones, 2007-08 Written for calculus-inclusive general chemistry courses, Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications. Unlike other texts, it begins with a detailed picture of the atom, then builds toward chemistry's frontier, continually demonstrating how to solve problems, think about nature and matter, and visualize chemical concepts as working chemists do. Flexibility in level is crucial and is largely established through clearly labeling, separating in boxes the calculus coverage in the text. Instructors have the option of whether to incorporate calculus in the coverage of topics. The multimedia integration of Chemical Principles is more deeply established than any other text for this course. Through the unique eBook, the comprehensive Chemistry Portal, Living Graph icons that connect the text to the Web, and a complete set of animations, students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding. *Prentice Hall Chemistry*, 2000 *Pennsylvania Game News*, 1947

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