

8. Calculate the pH corresponding to each of the following OH^- concentrations. Tell whether each solution is acidic, basic or neutral.

a. $[\text{OH}^-] = 1.00 \times 10^{-4} \text{ M}$

b. $[\text{OH}^-] = 4.56 \times 10^{-11} \text{ M}$

9. For each H^+ or OH^- concentration listed, calculate the concentration of the complementary ion AND the pH AND the pOH of the solution.

a. $[\text{OH}^-] = 4.29 \times 10^{-11} \text{ M}$

b. $[\text{H}^+] = 4.29 \times 10^{-11} \text{ M}$

10. Calculate the H^+ concentration for solutions with the following pH or pOH.

a. $\text{pH} = 5.99$

b. $\text{pOH} = 12.56$

11. What is the pH of a 0.075 M hydrobromic acid solution? *Hint: HBr is a STRONG acid*

12. Answer the following questions related to buffers:

a. What two components are required in every buffered solution?

b. Circle the pair shown below, that could act as a buffer in solution.



c. Which component will counteract the addition of an acid to the system? A base?

SKILL - The student will be able to calculate an acid - base stoichiometry problem.

13. Determine the concentration of potassium hydroxide that is required if 45.0 mL is necessary to neutralize 37.2 mL of 0.550 M sulfuric acid during a titration.

Acids & Bases: IN-CLASS Individual Skill Review

SKILL - The student will be able to identify the acid, base, conj acid & conj base in an acid - base reaction.

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