

Thales Jem Manual

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Thales Jem Manual:

Exploring Classical Greek Construction Problems with Interactive Geometry Software Ad Meskens, Paul Tytgat, 2017-02-02 In this book the classical Greek construction problems are explored in a didactical enquiry based fashion using Interactive Geometry Software IGS The book traces the history of these problems stating them in modern terminology By focusing on constructions and the use of IGS the reader is confronted with the same problems that ancient mathematicians once faced The reader can step into the footsteps of Euclid Vi te and Cusanus amongst others and then by experimenting and discovering geometric relationships far exceed their accomplishments Exploring these problems with the neusis method lets him discover a class of interesting curves By experimenting he will gain a deeper understanding of how mathematics is created More than 100 exercises guide him through methods which were developed to try and solve the problems The exercises are at the level of undergraduate students and only require knowledge of elementary Euclidean geometry and pre calculus algebra It is especially well suited for those students who are thinking of becoming a mathematics teacher and for mathematics teachers *The Athenaeum* ,1865 *Athenaeum* ,1865 **American Cinematographer** ,2007

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web oxidation reduction redox reactions oxidation $\text{Cu} \rightarrow \text{Cu}^{2+}$ reduction $\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$ when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} \rightarrow \text{MnO}_4^-$ is

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web 1 practice problem balance the redox reaction in basic solution $\text{NO}_2^- + \text{Al} \rightarrow \text{NH}_3 + \text{AlO}_2^-$ practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of OH^- $\text{Br}_2 + \text{I}^- \rightarrow \text{Br}^- + \text{I}_2$ previous topic next topic

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web google classroom for the following redox equation $\text{MnO}_4^- + \text{SO}_2 + \text{H}^+ \rightarrow \text{Mn}^{2+} + \text{HSO}_4^-$ what will be the stoichiometric coefficient of HSO_4^- if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{Au} + \text{CN}^- + 2\text{e}^-$ $2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ 2 make electrons equal
convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{Au} + \text{CN}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} \rightarrow \text{O}_2 + 2\text{H}^+$ 3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + 2\text{NO} + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow \text{TeO}_3^{2-} + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te}$ 10 $\text{ReO}_4^- + 11\text{PbO}_2 + \text{IO}_2^- + \text{Pb}^{2+} + 12\text{H}^+ \rightarrow \text{NO}_3^- + \text{IO}_3^- + \text{Re} + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b $\text{C}_6\text{H}_8\text{O}_6$ and $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 c $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 d $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 e $\text{C}_6\text{H}_8\text{O}_6$ and NO_2

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