

## Heat Capacity and Specific Heat Calculations - Practice

A 1.000 g sample of decane,  $C_{10}H_{22}$ , is ignited inside a bomb calorimeter. The temperature rises from  $20.0^{\circ}\text{C}$  to  $78.8^{\circ}\text{C}$ . Given that the heat capacity of the calorimeter is  $810.1 \text{ J}/^{\circ}\text{C}$ , determine the heat ( $q$ ) for the combustion of decane in units of  $\text{kJ/mol}$  of decane.

$$(1)(810.1)(58.8)$$

$$q = 47633.9 \text{ J}$$

Calculate the final temperature after  $1485 \text{ J}$  of energy is added to  $16.7 \text{ g}$  of water at  $23.4^{\circ}\text{C}$ .

$$q = mC\Delta T$$

$$1485 = (16.7)(4.184)(T_f - 23.4)$$

$$44.65^{\circ}\text{C}$$

$$1485 = (69.8728)(T_f) - 1635.024$$

$$3120.024 = 69.8728 T_f$$

How much heat is absorbed by a copper penny with a mass of  $3.10 \text{ g}$  when its temperature rises from  $-8.0^{\circ}\text{C}$  to  $37^{\circ}\text{C}$ ?

$$T_f = 37$$

$$T_i = -8.0$$

$$37 - (-8) = 45^{\circ}\text{C}$$

$$(3.10)(45)(0.385) = 53.71 \text{ J}$$

An unknown metal has a mass of  $19.7 \text{ g}$ . When  $8.20 \times 10^2 \text{ J}$  of heat are added to the sample, the sample temperature changes by  $55.5^{\circ}\text{C}$ . Calculate the specific heat capacity of the unknown metal.

$$mC\Delta T$$

$$m = 19.7$$

$$C =$$

$$\Delta T = 55.5$$

$$q = 820 \text{ J}$$

$$(19.7)(C)(55.5) = 820$$

$$C = 0.75 \text{ J/g}\cdot^{\circ}\text{C}$$

# Specific Heat Calculations With Answers

**Patrick Vollmar**



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