

Electron Configuration Practice
Chemistry

Name: _____

Due Date: _____

How to write an electron configuration:

- Determine the total number of electrons to be represented.
- Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- The sum of the superscripts should equal the total number of electrons. Example:
 $12\text{Mg } 1s^2 2s^2 2p^6 3s^2$

I. Configuration Writing Practice

- Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

1. Na: $1s^2 2s^2 2p^6 3s^1$

2. Pb: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

3. Sr: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$

4. U: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^3 6d^1$

5. N: $1s^2 2s^2 2p^3$

* 6. Ag: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^{10}$

7. Ti: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$

8. Ce: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^1 5d^1$

9. Cl: $1s^2 2s^2 2p^6 3s^2 3p^5$

10. Hg: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$

Skill Practice 13 Configuration Practice Answer Key

Chao Zhang



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Table of Contents Skill Practice 13 Configuration Practice Answer Key

1. Understanding the eBook Skill Practice 13 Configuration Practice Answer Key
 - The Rise of Digital Reading Skill Practice 13 Configuration Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 13 Configuration Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 13 Configuration Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 13 Configuration Practice Answer Key
 - Personalized Recommendations
 - Skill Practice 13 Configuration Practice Answer Key User Reviews and Ratings

- Skill Practice 13 Configuration Practice Answer Key and Bestseller Lists
- 5. Accessing Skill Practice 13 Configuration Practice Answer Key Free and Paid eBooks
 - Skill Practice 13 Configuration Practice Answer Key Public Domain eBooks
 - Skill Practice 13 Configuration Practice Answer Key eBook Subscription Services
 - Skill Practice 13 Configuration Practice Answer Key Budget-Friendly Options
- 6. Navigating Skill Practice 13 Configuration Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Skill Practice 13 Configuration Practice Answer Key Compatibility with Devices
 - Skill Practice 13 Configuration Practice Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 13 Configuration Practice Answer Key
 - Highlighting and Note-Taking Skill Practice 13 Configuration Practice Answer Key
 - Interactive Elements Skill Practice 13 Configuration Practice Answer Key
- 8. Staying Engaged with Skill Practice 13 Configuration Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 13 Configuration Practice Answer Key
- 9. Balancing eBooks and Physical Books Skill Practice 13 Configuration Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 13 Configuration Practice Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 13 Configuration Practice Answer Key
 - Setting Reading Goals Skill Practice 13 Configuration Practice Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 13 Configuration Practice Answer Key
 - Fact-Checking eBook Content of Skill Practice 13 Configuration Practice Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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