

SNC1D3 – Solubility Curve worksheet SOLUTIONS

For questions 1-5 the amount of solute is given, and a temperature is stated. If all the solute could be dissolved in 100 grams of water at the stated temperature, would the resulting solution be unsaturated, saturated, or supersaturated?

- 1) 60g KCl at 70°C **supersaturated [above line]**
- 2) 90g KNO₃ at 60°C **unsaturated [below line]**
- 3) 110g NaNO₃ at 45°C **saturated [on line]**
- 4) 10g KClO₃ at 10°C **supersaturated [above line]**
- 5) 60g NH₄Cl at 70°C **unsaturated [below line]**

For questions 6-10 a solute and a temperature are given. Tell How many grams of each solute must be added to 100 grams Of water to form a **saturated solution** (on the line) at that given temperature

- 6) NaNO₃ at 30°C **96g**
- 7) KClO₃ at 70°C **36g**
- 8) KNO₃ at 45°C **75g**
- 9) KCl at 40°C **59g**
- 10) NaCl at 90°C **40g**

For questions 11-13, tell which solution is more concentrated (more solute dissolved)

- 11) At 50°C - **A) a saturated solution of KNO₃ [85g]**
B) a saturated solution of NH₄Cl [52g]
- 12) At 50°C - **A) a saturated solution of KNO₃ [85g]**
B) an unsaturated solution of NaNO₃ consisting of 100g of the solute dissolved in 100g of water.
- 13) At 50°C - **A) a saturated solution of NaNO₃ [116g]**
B) a supersaturated solution of NH₄Cl consisting of 60g of the solute dissolved in 100g of water.

- 14) Which compound is most soluble at 40°C? **KI**
- 15) Which compound is least soluble at 90°C? **SO₂**
- 16) How many grams of KNO₃ will dissolve in 100g of water at 60°C? **108g**
- 17) If 100g of a saturated solution of KNO₃ at 70°C is cooled to 20°C, how much KNO₃ will crystallize (in grams) **100g**

- 18) If 200g of a saturated solution of KClO₃ at 50°C were evaporated to dryness, what would be left in the beaker?
42g KClO₃

OR $\frac{113.5\text{g}}{135\text{g}} = \frac{x}{21\text{g}} \Rightarrow 100x = 21 \times 200 \Rightarrow 100x = 4200 \Rightarrow x = \frac{4200}{100} \Rightarrow x = 42$

- 19) How much KCl will dissolve in 200g of water at 80°C **102g [51g x 2]**

- 20) Which is more soluble at 30°C? KNO₃ or NaNO₃? **NaNO₃**

- 21) How much NH₄Cl will dissolve in 50g of water at 60°C? **29g**

$\frac{58\text{g}}{100\text{g}} = \frac{x}{50\text{g}} \Rightarrow 100x = 50 \times 58 \Rightarrow 100x = 2900 \Rightarrow x = \frac{2900}{100} \Rightarrow x = 29$ OR $[58\text{g}/2 \text{ in } 100\text{g of water}/2 = 29]$

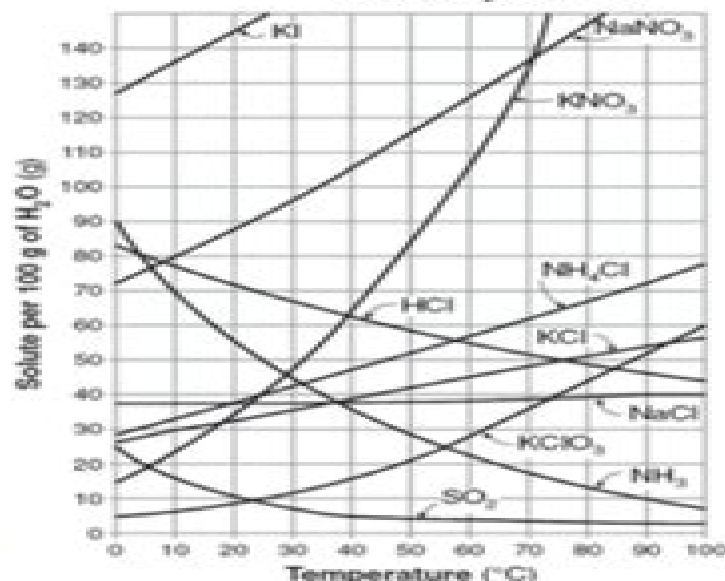
- 22) A saturated solution of NH₄Cl was made at 50°C in 100g of water. If the temperature is raised to 90°C, how much more NH₄Cl must be added to be saturated? **20g [52g in 50°C and 72g in 90°C $\Rightarrow 72-52 = 20$]**

- 23) A saturated solution of KClO₃ at 70°C in 200g of water is cooled to 18°C. How much solute will precipitate out? **56g [36g at 70°C in 100g of water so 36x2=72 at 70°C in 200g of water $\Rightarrow$ 8g at 18°C in 100g of water, so 8x2=16 at 18°C in 200g of water] [72 - 16 = 56]**

- 24) How much more SO₂ will dissolve in 50g of water at 10°C than at 40°C? **5.5g**

$[16/2 \text{ at } 10^\circ\text{C} = 8\text{g} \Rightarrow 5\text{g}/2 \text{ at } 40^\circ\text{C} = 2.5\text{g}] \quad [8 - 2.5 = 5.5]$

Solubility Curves



Solubility Curve Practice Problems Answer

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