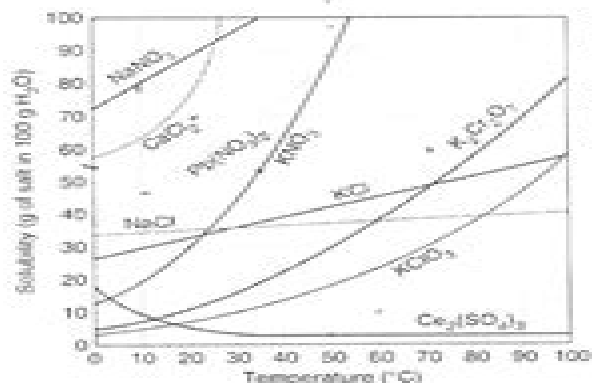


Worksheet: Solubility Graphs

Name

Key

Use the provided solubility graph to answer the following questions:



For questions 1 - 4 an amount of solute is given, and a temperature is stated. If all of the solute could be dissolved in 100 g of water at the given temperature, would the resulting solution be unsaturated, saturated, or supersaturated?

- 60 g KCl at 70 °C Super Saturated
- 10 g KClO₃ at 60 °C Unsaturated
- 80 g NaNO₃ at 10 °C Saturated
- 70 g CaCl₂ at 20 °C Unsaturated

For questions 5 - 8 a solute and temperature are given. Tell how many grams of each solute must be added to 100 g of water to form a saturated solution at the given temperature.

- Pb(NO₃)₂ at 10 °C 47g
- Ce₂(SO₄)₃ at 50 °C 3g
- NaCl at 20 °C 35g
- K₂Cr₂O₇ at 50 °C 30g

For questions 9 and 10 underline the solution that is more concentrated.

- At 10 °C: a saturated solution of KNO₃ or a saturated solution of CaCl₂
- At 50 °C: a saturated solution of KNO₃ or an unsaturated solution of NaNO₃ consisting of 90 g of the solute dissolved in 100 g of water.

For questions 11 - 12, show your work and circle your final answer.

- If 115 g KNO₃ are added to 100 g of water at 35 °C, how many grams do not dissolve? 55g = 60g will not dissolve.

- What mass of KCl would be needed to form a saturated solution if the KCl was dissolved in 200 g of water at 80 °C?

$$\frac{52g \text{ KCl}}{100g \text{ H}_2\text{O}} = \frac{104g \text{ KCl}}{200g \text{ H}_2\text{O}}$$

Same Ratio

Solubility Temperature Graphs Chapter 14 Worksheet Answers

Physical Science Study Committee

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the Physical Science Study Committee text.

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