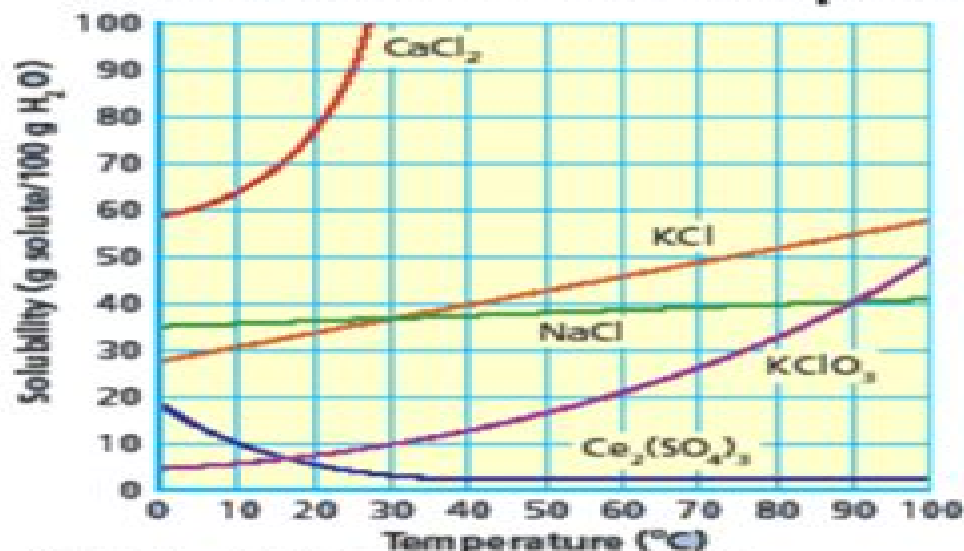


Directions: Use the graph below to answer the following questions. If the question requires a calculation, **SHOW ALL WORK TO RECEIVE CREDIT FOR THE ANSWER**. No work, no credit, no kidding! Remember units.

Solubilities as a Function of Temperature



Types of Solutions

- Points that collectively make up the solubility curve (points **ON** the curve) represent _____ solutions.
- Points that are **BELOW** the curve represent _____ solutions.
- Points **ABOVE** the solubility curve represent _____ solutions, and the difference between the point above the curve and on the curve represents the amount of solute which will precipitate out.

1. a. What is the solubility of calcium chloride at 5°C? _____ g/100g of H₂O
 b. What is the solubility of calcium chloride at 25°C? _____ g/100g of H₂O
 c. What is happening to the solubility and temperature?
 Solubility is _____ Increasing or Decreasing _____ as temperature is _____ Increasing or Decreasing _____.
 What is this type of relationship called? _____ direct or indirect _____.
 This occurs for which type of solutes? _____ gasses or solids _____.
2. a. What is the solubility of cerium sulfate at 10°C? _____ g/100g of H₂O
 b. What is the solubility of cerium sulfate at 50°C? _____ g/100g of H₂O
 c. What is happening to the solubility and temperature?
 Solubility is _____ increasing or decreasing _____ as temperature is _____ increasing or decreasing _____.
 What is this type of relationship called? _____ direct or indirect _____.
 Is this solution saturated, unsaturated, or supersaturated?
3. a. At 90°C, 10 g of potassium chlorate is dissolved in 100. g of water.
 Is this solution saturated, unsaturated, or supersaturated?
 b. How do you know?

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Karin Nielsen-Saines



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Table of Contents Solubility Practice Use The Graph On Answers

1. Understanding the eBook Solubility Practice Use The Graph On Answers
 - The Rise of Digital Reading Solubility Practice Use The Graph On Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Practice Use The Graph On Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Practice Use The Graph On Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Practice Use The Graph On Answers
 - Personalized Recommendations
 - Solubility Practice Use The Graph On Answers User Reviews and Ratings
 - Solubility Practice Use The Graph On Answers and Bestseller Lists
5. Accessing Solubility Practice Use The Graph On Answers Free and Paid eBooks

- Solubility Practice Use The Graph On Answers Public Domain eBooks
 - Solubility Practice Use The Graph On Answers eBook Subscription Services
 - Solubility Practice Use The Graph On Answers Budget-Friendly Options
6. Navigating Solubility Practice Use The Graph On Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Solubility Practice Use The Graph On Answers Compatibility with Devices
 - Solubility Practice Use The Graph On Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Solubility Practice Use The Graph On Answers
 - Highlighting and Note-Taking Solubility Practice Use The Graph On Answers
 - Interactive Elements Solubility Practice Use The Graph On Answers
8. Staying Engaged with Solubility Practice Use The Graph On Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Practice Use The Graph On Answers
9. Balancing eBooks and Physical Books Solubility Practice Use The Graph On Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Practice Use The Graph On Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solubility Practice Use The Graph On Answers
- Setting Reading Goals Solubility Practice Use The Graph On Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solubility Practice Use The Graph On Answers
- Fact-Checking eBook Content of Solubility Practice Use The Graph On Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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