

SCH4U1 Unit 5 - Thermochemistry and Rates of Reaction
Summative Assessment

Name: Answer keys

K/U:

17

TH:

10

C:

16

APP:

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1. Please answer the following questions in the space provided. [K5]

a) When solid KBr is dissolved in water, the solution gets colder. Is the chemical system exothermic or endothermic? [1]

endothermic

b) When a sample of potassium nitrate is dissolved in water, the temperature of the water changes from 23.5°C to 20.3°C. Is the chemical system exothermic or endothermic? [1]

endothermic

c) In the reaction $2A + B \rightarrow C + 3D$, reactant A is found to disappear at the rate of $6.2 \times 10^{-4} \text{ mol/L.s}$. What is the rate of disappearance of B? [1]

2 mol of A disappear at rate $6.2 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}$

1 mol of B disappears at $x \frac{\text{mol}}{\text{L.s}}$
 $x = \frac{6.2 \times 10^{-4} \text{ L.s}}{2} = 3.1 \times 10^{-4} \frac{\text{mol}}{\text{L.s}}$

d) The following diagram shows a collision between two BrNO molecules. Use collision theory to explain the results of the collision. [2]



Reaction did take place. New compounds formed: 2 NO and 1 Br₂ are the products. It was an effective collision. Reactants collided with proper orientation and sufficient (optimal) kinetic energy.

2. Consider the following reaction:

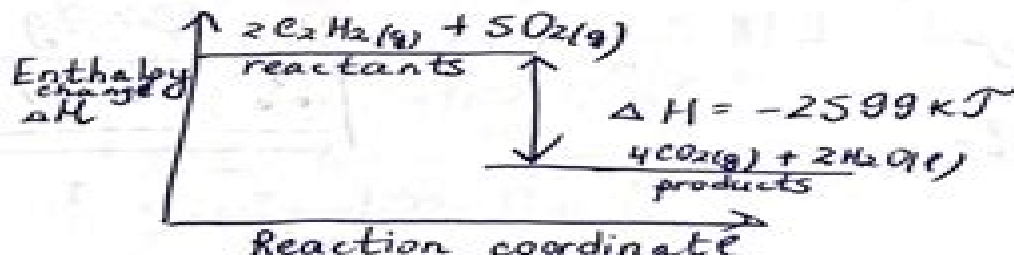


$$\Delta H = -2599 \text{ kJ}$$

a) Write a thermochemical equation with the enthalpy change, ΔH , 'inside' the equation. [C1]



b) Sketch an enthalpy diagram for this reaction. Provide appropriate labels for each axis. [C3]



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Terry C. Jones



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