

LAB 13: SOLUBILITY PRODUCT CONSTANT



Goals

1. Learn of solubility product constants
2. Estimate a solubility constant
3. Observe that a solubility constant is a special case of an equilibrium constant

Materials and Equipment

Reagents:

0.1 M calcium nitrate, $\text{Ca}(\text{NO}_3)_2$

0.1 M sodium oxalate, $\text{Na}_2\text{C}_2\text{O}_4$

Pipet, mini

Toothpick, plastic

Reaction plate, 96 well

Materials Not Included

A sheet of black paper

Distilled water

Introduction

Ionic compounds, such as salts, dissociate to form ions when they dissolve in water. An example is sodium chloride:



Solid NaCl dissociates to become Na^+ and Cl^- ions. The Na^+ cations are attracted to the negative "ends" of water molecules, and the Cl^- anions are attracted to the positive "ends" of water molecules.

In other compounds the attraction between the cations and the anions is stronger, and there is little dissociation. These are sparingly or slightly soluble salts. An example is silver chloride, AgCl . A beaker of this solution often has solid AgCl crystals in the bottom. The crystals are

in equilibrium with the Ag^+ and Cl^- ions that are in the solution:



In this reaction, the rate of AgCl dissolving is equal to the rate of Ag^+ and Cl^- ions combining to form solid AgCl . In other words, the reaction is going both ways at the same time and at the same rate. It is said to be in equilibrium. The equilibrium constant is:

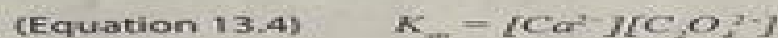


$\text{AgCl}(s)$ is not considered in the equilibrium constant because it is a solid.

We call this special equilibrium constant the solubility product constant, K_{sp} .

Some of the beautiful structures you may have seen in caves are composed of sparingly soluble calcium carbonate. This material slightly dissolves in water, and later precipitates when the water in the solution evaporates or the solution cools.

Another sparingly soluble salt is calcium oxalate, which sometimes precipitates in our kidneys as kidney stones. These can be very painful and often require medical treatment. We will study calcium oxalate, CaC_2O_4 , in this experiment. The solubility product constant for this is:



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Table of Contents Solubility Lab Answer Key

1. Understanding the eBook Solubility Lab Answer Key
 - The Rise of Digital Reading Solubility Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Lab Answer Key
 - Personalized Recommendations
 - Solubility Lab Answer Key User Reviews and Ratings

- Solubility Lab Answer Key and Bestseller Lists
- 5. Accessing Solubility Lab Answer Key Free and Paid eBooks
 - Solubility Lab Answer Key Public Domain eBooks
 - Solubility Lab Answer Key eBook Subscription Services
 - Solubility Lab Answer Key Budget-Friendly Options
- 6. Navigating Solubility Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Solubility Lab Answer Key Compatibility with Devices
 - Solubility Lab Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solubility Lab Answer Key
 - Highlighting and Note-Taking Solubility Lab Answer Key
 - Interactive Elements Solubility Lab Answer Key
- 8. Staying Engaged with Solubility Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Lab Answer Key
- 9. Balancing eBooks and Physical Books Solubility Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Lab Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solubility Lab Answer Key
 - Setting Reading Goals Solubility Lab Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solubility Lab Answer Key
 - Fact-Checking eBook Content of Solubility Lab Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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