

# Table F

## Solubility Guidelines for Aqueous solutions

| Ions That Form Soluble Compounds                                            | Exceptions                                                                                                      | Ions That Form Insoluble Compounds* | Exceptions                                                                                                                |
|-----------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------|-------------------------------------|---------------------------------------------------------------------------------------------------------------------------|
| Group 1 ions ( $\text{Li}^+$ , $\text{Na}^+$ , etc.)                        |                                                                                                                 | carbonate ( $\text{CO}_3^{2-}$ )    | when combined with Group 1 ions or ammonium ( $\text{NH}_4^+$ )                                                           |
| ammonium ( $\text{NH}_4^+$ )                                                |                                                                                                                 | chromate ( $\text{CrO}_4^{2-}$ )    | when combined with Group 1 ions, $\text{Ca}^{2+}$ , $\text{Mg}^{2+}$ , or ammonium ( $\text{NH}_4^+$ )                    |
| nitrate ( $\text{NO}_3^-$ )                                                 |                                                                                                                 | phosphate ( $\text{PO}_4^{3-}$ )    | when combined with Group 1 ions or ammonium ( $\text{NH}_4^+$ )                                                           |
| acetate ( $\text{C}_2\text{H}_3\text{O}_2^-$ or $\text{CH}_3\text{COO}^-$ ) |                                                                                                                 | sulfide ( $\text{S}^{2-}$ )         | when combined with Group 1 ions or ammonium ( $\text{NH}_4^+$ )                                                           |
| hydrogen carbonate ( $\text{HCO}_3^-$ )                                     |                                                                                                                 | hydroxide ( $\text{OH}^-$ )         | when combined with Group 1 ions, $\text{Ca}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Sr}^{2+}$ , or ammonium ( $\text{NH}_4^+$ ) |
| chlorate ( $\text{ClO}_3^-$ )                                               |                                                                                                                 |                                     |                                                                                                                           |
| halides ( $\text{Cl}^-$ , $\text{Br}^-$ , $\text{I}^-$ )                    | when combined with $\text{Ag}^+$ , $\text{Pb}^{2+}$ , or $\text{Hg}_2^{2+}$                                     |                                     |                                                                                                                           |
| sulfates ( $\text{SO}_4^{2-}$ )                                             | when combined with $\text{Ag}^+$ , $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , or $\text{Pb}^{2+}$ |                                     |                                                                                                                           |

\*compounds having very low solubility in  $\text{H}_2\text{O}$

# Solubility Guidelines For Aqueous Solutions Answers

**The Princeton Review**



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