

Binary Ionic Compounds

Directions: First quickly scan the worksheet and circle any metals (as a symbol or as a name) that is a transition metal. Then either give the same or chemical formula as necessary for each problem below.

Remember that transition metals can have multiple oxidation states, so you are required to indicate the appropriate oxidation state in parenthesis when naming them. You must use the subscript and oxidation state of the anion to help determine the oxidation state of the transition metal.

- | | |
|---|---|
| 1. LiF <u>lithium fluoride</u> | 2. lithium chloride <u>LiCl</u> |
| 3. Li ₂ O <u>lithium oxide</u> | 4. lithium nitride <u>Li₃N</u> |
| 5. Li ₃ P <u>lithium Phosphide</u> | 6. beryllium fluoride <u>BeF₂</u> |
| 7. BeO <u>Beryllium Oxide</u> | 8. beryllium sulfide <u>BeS</u> |
| 9. BF ₃ <u>Boron fluoride Fluoride</u> | 10. boron chloride <u>BCl₃</u> |
| 11. BBr ₃ <u>Boron Bromide</u> | 12. boron oxide <u>B₂O₃</u> |
| 13. BN <u>Boron nitride</u> | 14. sodium fluoride <u>NaF</u> |
| 15. CuF <u>Copper (I) Fluoride</u> | 16. copper (I) chloride <u>CuCl</u> |
| 17. Cu ₂ O <u>Copper (I) oxide</u> | 18. copper (II) oxide <u>Cu₂O</u> |
| 19. Cu ₃ N <u>Copper (I) nitride</u> | 20. copper (II) nitride <u>Cu₃N</u> |
| 21. Na ₂ O <u>Sodium Oxide</u> | 22. sodium nitride <u>Na₃N</u> |
| 23. PbF ₂ <u>lead (II) Fluoride</u> | 24. lead (IV) fluoride <u>PbF₄</u> |
| 25. PbS <u>lead (II) sulfide</u> | 26. lead (IV) sulfide <u>PbS₂</u> |
| 27. Pb ₃ N ₄ <u>lead (II) nitride</u> | 28. lead (IV) oxide <u>PbO₂</u> |
| 29. FeF ₃ <u>iron (III) fluoride</u> | 30. iron (III) bromide <u>FeBr₃</u> |
| 31. Fe ₂ O ₃ <u>iron (III) oxide</u> | 32. chromium (VI) phosphide <u>CrP₂</u> |
| 33. FeP <u>iron (III) Phosphide</u> | 34. aluminum fluoride <u>AlF₃</u> |
| 35. Al ₂ <u>Aluminum iodide</u> | 36. aluminum oxide <u>Al₂O₃</u> |
| 37. Mn ₂ O ₇ <u>Manganese (VII) oxide</u> | 38. manganese (IV) bromide <u>MnBr₄</u> |
| 39. SnSe <u>tin (II) selenide</u> | 40. tin (IV) oxide <u>SnO₂</u> |

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Table of Contents Simple Binary Ionic Compounds Answer Key

1. Understanding the eBook Simple Binary Ionic Compounds Answer Key
 - The Rise of Digital Reading Simple Binary Ionic Compounds Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Simple Binary Ionic Compounds Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Simple Binary Ionic Compounds Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Simple Binary Ionic Compounds Answer Key
 - Personalized Recommendations

- Simple Binary Ionic Compounds Answer Key User Reviews and Ratings
- Simple Binary Ionic Compounds Answer Key and Bestseller Lists
- 5. Accessing Simple Binary Ionic Compounds Answer Key Free and Paid eBooks
 - Simple Binary Ionic Compounds Answer Key Public Domain eBooks
 - Simple Binary Ionic Compounds Answer Key eBook Subscription Services
 - Simple Binary Ionic Compounds Answer Key Budget-Friendly Options
- 6. Navigating Simple Binary Ionic Compounds Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Simple Binary Ionic Compounds Answer Key Compatibility with Devices
 - Simple Binary Ionic Compounds Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Simple Binary Ionic Compounds Answer Key
 - Highlighting and Note-Taking Simple Binary Ionic Compounds Answer Key
 - Interactive Elements Simple Binary Ionic Compounds Answer Key
- 8. Staying Engaged with Simple Binary Ionic Compounds Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Simple Binary Ionic Compounds Answer Key
- 9. Balancing eBooks and Physical Books Simple Binary Ionic Compounds Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Simple Binary Ionic Compounds Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Simple Binary Ionic Compounds Answer Key
 - Setting Reading Goals Simple Binary Ionic Compounds Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Simple Binary Ionic Compounds Answer Key
 - Fact-Checking eBook Content of Simple Binary Ionic Compounds Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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