

Name: _____

Date: _____

Significant Figures

Determine the number of significant figures in each of these numbers.

	Number	Significant Figures		Number	Significant Figures
1.	357	_____	2.	10000	_____
3.	51015	_____	4.	6.060×10^{-2}	_____
5.	0.0007	_____	6.	4.556×10^{-9}	_____
7.	5050	_____	8.	5050.0	_____
9.	6.8×10^3	_____	10.	0.002110	_____
11.	33.303	_____	12.	170	_____
13.	15.0×10^{-5}	_____	14.	0.7007	_____
15.	0.70070	_____	16.	4206	_____
17.	0.02	_____	18.	10.01	_____
19.	0	_____	20.	0.0	_____

Significant Figures Worksheet Answer Key Chemistry 1

Brendan G. Carr



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and computational chemistry used to interpret laser experiments and the applications of lasers in spectroscopy and photochemistry Experiments range from simple in class demonstrations to more elaborate configurations for advanced students Each chapter has historical and theoretical background as well as options suggested for variations on the prescribed experiments This text will be useful for undergraduate students in advanced lab classes for instructors designing these classes or for graduate students beginning a career in laser science It can also be used as a supplementary text for courses in molecular spectroscopy or optics

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FL Chemistry 1 - Worksheet Collection Lincoln Learning Solutions,2024-12-09 **The Answer Key: A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success (Volume One) (First Edition)** Rachel Turoscy,2018-08-09 The Answer Key A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success Volume 1 is a concise and accessible textbook that covers the critical information a student needs to understand the basic mathematics used in chemistry courses The book provides easy to understand step by step instructions for solving general chemistry problems The book begins with chapters dedicated to problem solving methodology and unit conversions In subsequent chapters the text covers important topics like ionic and covalent bonding chemical formula calculations solubility and reactions in aqueous solution gases the first law of Thermodynamics Quantum theory and electron configuration It also covers periodic trends the Lewis Dot Structures and bonding theories Each chapter contains sample problems and practice problems to help further understanding of how math and chemistry go hand in hand The Answer Key is an excellent resource for any undergraduate course that deals with the basic concepts of general chemistry

Chemistry, Problem-Solving Worktext James E. Brady,Frederick Senese,2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the

chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

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