

Names: \_\_\_\_\_

## Thermodynamics IN CLASS Skill Review

**Skill -** Can define and identify direction of heat flow for exothermic and endothermic situations. Before completing the calculations below, answer the following:

- Provide the labels exothermic and endothermic, to the following pictures:

a. What are some everyday examples of

**endothermic:**

exothermic and endothermic reactions?

Answers may vary - exothermic (hot pad),

endothermic (baking a cake).

**exothermic:**



- Label the following reactions as being exothermic or endothermic:

a. As a tree grows, photosynthesis **absorbs** energy from the environment to break apart  $\text{CO}_2$  and  $\text{H}_2\text{O}$ .

**endothermic**

b. Sweating cools a person down as energy is **released** from the skin to change water into a gas.

**exothermic**

**Skill -** Use the equation  $\Delta E = q + w$  correctly when given these terms in a real-world situation.

For the following, calculate the  $\Delta E$  and label each scenario as describing an exothermic or endothermic reaction:

- A system releases 23 J of heat while 12 J of work is done on it. What is the  $\Delta E$ ?

$$\Delta E = q + w \rightarrow -23\text{J} + 12\text{J} = \mathbf{-11\text{J} \text{ exothermic}}$$

- If a system absorbs 67 J of heat and performs 32 J of work, what is the  $\Delta E$ ?

$$\Delta E = q + w \rightarrow 67\text{J} + -32\text{J} = \mathbf{35\text{J} \text{ endothermic}}$$

**Skill -** Use  $q = m \cdot c \cdot \Delta T$  to calculate the missing quantity, given the others.

Be sure to show all work required for the following calculations:

- Graphite has specific heat of  $0.709 \text{ J/g}^\circ\text{C}$ . If a 25.0g piece of graphite is cooled from  $35^\circ\text{C}$  to  $18^\circ\text{C}$ , how much energy was lost by the graphite? Is the reaction endothermic or exothermic?

$$q = m \cdot c \cdot \Delta T \rightarrow 0.709\text{J/g}^\circ\text{C} \times 25.0\text{g} \times -17^\circ\text{C} = \mathbf{-300\text{J} \text{ exothermic}}$$

- An unknown substance with a mass of 100.0g, **absorbs**  $1.0 \times 10^3 \text{ J}$  while undergoing a temperature increase of  $25^\circ\text{C}$ . What is the specific heat of the substance? Is the reaction endothermic or exothermic?

$$q = m \cdot c \cdot \Delta T \rightarrow c = q / (m \cdot \Delta T) \rightarrow 1.0 \times 10^3 \text{ J} / (100.0\text{g} \times 25^\circ\text{C}) = \mathbf{0.4\text{J/g}^\circ\text{C}} \text{, } q \text{ is endothermic}$$

# Smurray Electricity And Thermodynamics Review

## Answer Key

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