



Slow Cooker Recipe For Pork Tenderloin

O García



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$$

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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation
$$2\text{H}_2\text{S} + 3\text{O}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{SO}_2$$

8 3 mole $\text{H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149 4 \text{ g H}_2\text{O}$ is your theoretical yield so percent yield
$$\frac{137 1 \text{ g H}_2\text{O}}{149 4 \text{ g H}_2\text{O}} \times 100 = 91 77\%$$

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO}_2 + 3\text{H}_2\text{O}$$

if the reaction of 21 3 grams of O_2 produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have
$$\frac{21 3 \text{ g O}_2}{32 \text{ g O}_2/\text{mole}} = 0 666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained
$$\frac{\text{calculated mass of the product}}{\text{mass of the product expected}} \times 100$$

in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x 1 mol Al x 2 mol Al_2O_3 x 101.964 g Al_2O_3 302.3 g Al_2O_3 26.982 g Al 4 mol Al 1 mol Al_2O_3 c determine the percent yield $\frac{\text{yield actual} \times 100}{260.0 \text{ g} \times 100} = \frac{86.01}{302.3} \times 100$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage
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