CHEMICAL EQUATIONS

A chemical equation is a simple and easy way to describe a chemical reaction. Chemical equations use chemical formulas and other symbols instead of words to summarize a reaction. A chemical equation tells you the substances you start with in a reaction and the substances that are formed in the end. The chemical equation for the formation of water is shown below.



All chemical equations share a common structure. The substances you have at the beginning are called the **reactants**. The new substances formed at the end of the reaction are called the **products**. The formulas for the reactants are written on the left side of the equation, followed by an **arrow** (->) which means "turns into" or "yields". The formulas for the products are written on the right side of the equation. When there are two or more reactants or products, they are separated by plus signs.

The law of conservation of mass was first demonstrated by a scientist named Antoine Lavoisier in the late 1700s. The law of conservation of mass states that in a chemical reaction, the total mass of the reactants must equal the total mass of the products. In an open system, matter can enter or escape from the surroundings. For example, a match burning in the open air is an open system. You cannot measure the mass of all the reactants and products in an open system. In a closed system, matter cannot enter or escape into the surroundings. A match burning inside a sealed glass chamber is an example of a closed system. A closed system allows you to measure the mass of all the reactants and products in a reaction.

Because of the law of conservation of mass, a chemical equation must be **balanced**. An equation is balanced when the total number of atoms of the reactants equals the total number of atoms in the products. To balance a chemical reaction, you may have to use coefficients. A coefficient is a number placed in front of a chemical formula in an equation. The coefficient tells you how many molecules of a reactant or product take part in a reaction.



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- 1. Understanding the eBook Webquest Chemical Formulas And Equations
 - The Rise of Digital Reading Webquest Chemical Formulas And Equations
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Webquest Chemical Formulas And Equations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Webquest Chemical Formulas And Equations
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Webquest Chemical Formulas And Equations
 - Personalized Recommendations

- Webquest Chemical Formulas And Equations User Reviews and Ratings
- Webguest Chemical Formulas And Equations and Bestseller Lists
- 5. Accessing Webquest Chemical Formulas And Equations Free and Paid eBooks
 - Webquest Chemical Formulas And Equations Public Domain eBooks
 - Webquest Chemical Formulas And Equations eBook Subscription Services
 - Webquest Chemical Formulas And Equations Budget-Friendly Options
- 6. Navigating Webquest Chemical Formulas And Equations eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Webquest Chemical Formulas And Equations Compatibility with Devices
 - Webquest Chemical Formulas And Equations Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Webquest Chemical Formulas And Equations
 - Highlighting and Note-Taking Webquest Chemical Formulas And Equations
 - Interactive Elements Webquest Chemical Formulas And Equations
- 8. Staying Engaged with Webquest Chemical Formulas And Equations
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Webquest Chemical Formulas And Equations
- 9. Balancing eBooks and Physical Books Webguest Chemical Formulas And Equations
 - ∘ Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Webquest Chemical Formulas And Equations
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Webquest Chemical Formulas And Equations
 - Setting Reading Goals Webquest Chemical Formulas And Equations
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Webquest Chemical Formulas And Equations
 - Fact-Checking eBook Content of Webquest Chemical Formulas And Equations

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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