

2-4: Precipitation Reactions

1. Start *Virtual ChemLab*, select *Reactions and Stoichiometry*, and then select *Precipitation Reactions* from the list of assignments. The lab will open in the Inorganic laboratory.
2. React each of the cations (across the top) with each of the anions (down the left) according to the data table using the following procedures:

Data Table

	AgNO_3 (Ag^+)	$\text{Pb}(\text{NO}_3)_2$ (Pb^{2+})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{2+})
Na_2CO_3 (CO_3^{2-})	<i>a</i>	<i>f</i>	<i>k</i>
Na_2S (S^{2-})	<i>b</i>	<i>g</i>	<i>l</i>
NaOH (OH^-)	<i>c</i>	<i>h</i>	<i>m</i>
Na_2SO_4 (SO_4^{2-})	<i>d</i>	<i>i</i>	<i>n</i>
NaCl (Cl^-)	<i>e</i>	<i>j</i>	<i>o</i>

- a. Drag a test tube from the box and place it on the metal test tube stand. You can then click on the bottle of Ag^+ ion solution on the shelf to add it to the test tube.
 - b. Click on the **Divide** button on the edge of the lab bench four times to make four additional test tubes containing Ag^+ . With one test tube in the metal stand and four others in the blue rack, click on the Na_2CO_3 bottle on the reagent shelf and observe what happens in the window at the top left. Record your observation in the table above. If the solution remains clear, record NR for no reaction. Drag this test tube to the red disposal bucket on the right.
 - c. Drag a second tube from the blue rack to the metal stand. Add Na_2S , record your observations and discard the tube. Continue with the third, fourth and fifth tube, but add NaOH , Na_2SO_4 , and NaCl respectively. Record your observations and discard the tubes. When you are finished, click on the red disposal bucket to clear the lab.
 - d. Repeat steps a-c for five test tubes of Pb^{2+} and Ca^{2+} . Record your observations in the data table. If no precipitate forms write NR for no reaction.
3. What happens in grid space *d*? What other reactions give similar results? Is it necessary to write an equation when no reaction occurs? Explain. _____

 4. Write balanced equations for all precipitation reactions you observed.
 - a.
 - b.
 - c.

Virtual Lab Precipitation Reactions Answers

KJ Lindholm-Leary



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Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001 **Make the Grade in AS and A2 Chemistry** George Facer, 2003 Make the Grade offers comprehensive exam support for AS and A2 Chemistry Part of the Nelson Advanced Science series it provides activities and questions for use throughout the course with exam questions including synoptic questions to help students fully prepare for examinations Acid Precipitation ,1986

Green Chemistry Laboratory Manual for General Chemistry Sally A. Henrie, 2015-03-18 Green chemistry involves designing novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts Providing educational laboratory materials that challenge students with the customary topics found in a general chemistry laboratory manual this lab manual enables students to see how green chemistry principles can be applied to real world issues Following a consistent format each lab experiment includes objectives prelab questions and detailed step by step procedures for performing the experiments Additional questions encourage further research about how green chemistry principles compare with traditional more hazardous experimental methods Cumulated Index Medicus ,1964 **U.S. Government Research Reports** ,1959 **Crime Laboratory Digest** ,1990 **Methods in**

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