# Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use q = (m)(Cp)(AT) to solve the following problems. Show all work and units.

How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$q = ?$$
  $\Delta t = 55 - 22$   $q = m \cdot \Delta t$   
 $m = 10$   $q = \frac{33}{2}$   $q = m \cdot \Delta t$   $q = 10 \times 0$ 

The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did 1. this sample absorb? C for water = 4.18 J/g°C

25 How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$bt = 25 - 10 = 15$$
 $m = 2259$ 
 $S = 4.18 5/5c$ 

$$a_1 = m \le 8t$$
  
 $= 22.5 \times 4.17 \times 15$   
 $= 1410755$   
 $IJ = 1 \times 10^3 \text{ KS}$   $= 14.1075 \text{ KS}$ 

To bring 1.0kg of water from 25°C to 99°C takes how much heat input? m= 1 14= 10000 91-1000 X 4-18 X 74

An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal?

A calorimeter has a heat capacity of 1265 J/C. A reaction causes the temperature of the 5. calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process? C = 1265 5/c.

$$C = 1265 \text{ J/c}$$

$$\Delta t = 25.12 - 22.34 = 2.78$$

$$Q = ?$$

$$Q = C\Delta t$$

$$= 1265 \times 2.78$$

$$= 3516.75$$

# **Using Specific Heat Answer Key**

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#### **Using Specific Heat Answer Key:**

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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Using Specific Heat Answer Key**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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