

Name: KEY

Section: _____

Chapter 7 Practice Worksheet: Covalent Bonds and Molecular Structure

1) How are ionic bonds and covalent bonds different?

Ionic bonds result from the transfer of electrons from one atom to another; Covalent bonds result from two atoms sharing electrons.

2) Describe the relationship between the length of a bond and the strength of that bond.

Strength of a bond increases as the bond gets shorter (inverse relationship)

3) Identify the type(s) of bond(s) found in the following molecules:

- a. CCl_4 covalent
- b. Li_2O ionic
- c. NF_3 covalent
- d. CaSO_4 ionic and covalent
- e. SO_2 covalent
- f. $\text{Mg}(\text{OH})_2$ ionic and covalent

4) Define electronegativity.

Electronegativity is the ability for an atom in a molecule to attract electrons to itself.

5) Use electronegativity values to place the following elements in **increasing** order: F, N, H, P, Si, C, O

$\text{Si} < \text{C} \approx \text{P} \approx \text{H} < \text{N} < \text{O} < \text{F}$

6) Determine if the bond between atoms in each example below is nonpolar covalent, polar covalent, or ionic.

- | | | | |
|------------------|--------------|-----------------|--------------|
| a. H_2 | <u>npc</u> | e. NF | <u>pc</u> |
| b. PCl | <u>pc</u> | f. MgO | <u>ionic</u> |
| c. F_2 | <u>npc</u> | g. CH | <u>npc</u> |
| d. NaBr | <u>ionic</u> | h. HCl | <u>pc</u> |

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