

Final mass of copper : 13.103 g of Cu

(mass evaporating dish + copper - mass evaporating dish)

$$67.245\text{g} - 54.142\text{g} = 13.103\text{g}$$

Final moles of copper : 0.206 moles of Cu

(mass in grams divided by 1 mole, multiplied by mass of element)

$$\frac{13.103\text{g Cu}}{63.55\text{g Cu}} = 0.206\text{ moles Cu}$$

Initial moles of copper : 0.032 moles of Cu

(mass in grams divided by 1 mole, multiplied by mass of element)

$$\frac{2.03\text{g Cu}}{63.55\text{g Cu}} = 0.032$$

Percent yield : 64.5%

$$\left(\frac{\text{actual yield}}{\text{theoretical yield}} \times 100\% \right) \text{ or in simpler terms for this lab... } \left(\frac{\text{recovered mass}}{\text{initial mass}} \times 100\% \right)$$

$$\frac{13.103}{2.03} \times 100 = 64.5\%$$

Stoichiometry Using Copper Lab 3answer Key

CL Gary



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Laboratory Manual to Accompany Chemistry, [by] Stanley R. Radel, Marjorie H. Navidi Arthur D. Baker, 1990

Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 provides a complete guide to the fundamentals of chemistry Page 4 of cover

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