#### Balance the equation below

$$1 CH_4 + 2 O_2 \rightarrow 1 CO_2 + 2 H_2O$$

27) Once the equation is balanced, what is the molar ratio of water (H<sub>2</sub>O) to oxygen (O<sub>2</sub>)? \_\_\_\_\_\_1:1

28) If 4.59 grams of O<sub>2</sub> react, how many grams of CH<sub>4</sub> are needed?

29) How many grams of CO<sub>2</sub> are needed to react with 8.95 grams of H<sub>2</sub>O?

Answer: 10.93 g CO:

#### Balance the equation below

$$2 H_2 + 1 O_2 \rightarrow 2 H_2O$$

30) Once the equation is balanced, what is the molar ratio of H<sub>2</sub> to O<sub>2</sub> ? \_\_\_\_\_\_\_ 2:1

31) How many moles of H<sub>2</sub> are needed to react with 0.42 moles of H<sub>2</sub>O?

Answer: 0.42 moles H<sub>2</sub>

32) How many grams of H<sub>2</sub>O are needed to react with 4.55 moles of H<sub>2</sub>?

33) How many moles of  $H_2O$  are produced when reacted with 8.35 grams of  $H_2$ ?

# **Stoichiometry Study Guide Chemistry Answer Key**

**Arshad Iqbal** 

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