

CHAPTER 3

STOICHIOMETRY

Questions

25. The atomic mass of any particular isotope is a relative mass to a specific standard. The standard is one atom of the carbon-12 isotope weighing exactly 12.0000 u. One can determine from experiment how much heavier or lighter any specific isotope is than ^{12}C . From this information, we assign an atomic mass value to that isotope. For example, experiment tells one that ^{16}O is about 4/3 heavier than ^{12}C , so a mass of $4/3(12.00) = 16.00$ u is assigned to ^{16}O .

The atomic mass listed in the periodic table is also an average mass. Most elements in nature occur as a mixture of isotopes. The atomic mass of an element is the average mass of all the isotopes that make up a specific element, weighted by abundance.

26. Consider a sample of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$. The molar mass of glucose is 180.16 g/mol. The chemical formula allows one to convert from molecules of glucose to atoms of carbon, hydrogen, or oxygen present and vice versa. The chemical formula also gives the mole relationship in the formula. One mole of glucose contains 6 mol C, 12 mol H, and 6 mol O. Thus mole conversions between molecules and atoms are possible using the chemical formula. The molar mass allows one to convert between mass and moles of compound, and Avogadro's number (6.022×10^{23}) allows one to convert between moles of compound and number of molecules.

27. Avogadro's number of dollars = 6.022×10^{23} dollars/mol dollars

$$\frac{1 \text{ mol dollars} \times \frac{6.022 \times 10^{23} \text{ dollars}}{\text{mol dollars}}}{7 \times 10^8 \text{ people}} = 8.6 \times 10^{13} = 9 \times 10^{13} \text{ dollars/person}$$

1 trillion = 1,000,000,000,000 = 1×10^{12} ; each person would have 90 trillion dollars.

28. Molar mass of CO_2 = $12.01 + 2(16.00) = 44.01$ g/mol

One mol of CO_2 contains 6.022×10^{23} molecules of CO_2 , 6.022×10^{23} atoms of C, and 1.204×10^{24} atoms of O. We could also break down 1 mol of CO_2 into the number of protons and the number of electrons present (1.325×10^{23} protons and 1.325×10^{23} electrons). In order to determine the number of neutrons present, we would need to know the isotope abundances for carbon and oxygen.

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