_____ Date: _____ Class period: _____ Name:

Unit 9: Stoichiometry Test Review

Directions: Please use dimensional analysis for the following calculations. Include the correct significant figures, substance and units.

 How many moles of calcium nitrate would react with 4.55 moles of chromium (III) sulfate to produce calcium sulfate and chromium(III) nitrate?

 $Ca(NO_1)_2 + Cr_2(SO_4)_1 \rightarrow CaSO_4 + Cr(NO_1)_3$

2. What mass of gallium chloride would react with excess fluorine to produce 78.9 g of chlorine? Gallium fluoride is the other product.

__ GaCl₃ + __ F₂ -> __ GaF₃ + __ Cl₂

3. How many grams of ammonium carbonate are needed to decompose in order to produce 6.52g of carbon dioxide? Ammonium oxide is the other product.

 $(NH_4)_2CO_3 \rightarrow (NH_4)_2O_5 + CO_2$

4. What volume of nitrogen gas at STP would react with 37.2 g of magnesium to produce magnesium nitride?

 $_$ Mg + $_$ N₂ \rightarrow $_$ Mg₁N₂

5. How many molecules of sulfur trioxide would react with 68.9 L of water at STP to produce sulfuric acid?

___ SO₃ + ___ H₂O → ___ H₂SO₄

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Paul Foglino, Princeton Review (Firm)

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