- Remember, this is practice Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice
  without a calculator, even if it slows you down. Look for the "easy math" common factors and rough estimation do
  not do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.
- 1. Balance the following equation:

 $NH_1 + O_2 \rightarrow NO_2 + H_2O$ 

The balanced equation shows that 1.00 mole of NHs requires \_\_\_ mole(s) of O<sub>2</sub>

- 0.57
- b. 1.25
- c. 1.33
- d. 1.75
- e. 3.5
- Write a balanced equation for the combustion of acetaldehyde, CH<sub>2</sub>CHO.

When properly balanced, the equation indicates that mole(s) of O<sub>2</sub> are required for each mole of CH<sub>2</sub>CHO.

- m. 1
- bc 2:
- c. 2.5
- d. 3
- e. 5
- What is the total mass of products formed when 16 grams of CH<sub>4</sub> is burned with excess oxygen?
  - in. 32 m.
  - b. 36 gr
  - c. 44 g.
  - d. 62 g
  - c. 80 g
- Write a balanced equation for the combustion of propane, C<sub>3</sub>H<sub>8</sub>.

When balanced, the equation indicates that \_\_\_\_ moles of O<sub>2</sub> are required for each mole of C<sub>3</sub>H<sub>6</sub>.

- m. 1.5
- ъ. з.
- c. 3.5
- dl. 5
- er. 80

 Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the <u>sum</u> of the coefficients in the balanced equation:

 $KClO_3 \rightarrow KCl + O_2$ 

- a. 3
- b. 5
- c. 6
- d. 7
- e. 8
- Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below.

- a. 1.5 g
- b. 2.0 g
- c. 3.0 g
- d. 6.0 g.
- c. 12 g;
- How many grams of nitric acid, HNO<sub>3</sub>, can be prepared from the reaction of 138 g of NO<sub>2</sub> with 54.0 g H<sub>2</sub>O according to the equation below?

$$3NO_2 + H_2O \rightarrow 2HNO_1 + NO$$

- m. 92
- 6 108.
- c. 126
- d. 189
- e. 279
- 8. Which of the following statements is true?
  - The molar mass of CaCO<sub>3</sub> is 100.1 g mol<sup>-1</sup>.
  - 50 g of CaCO<sub>3</sub> contains 9 × 10<sup>23</sup> oxygen atoms.
  - III. A 200 g sample of CaCO<sub>1</sub> contains 2 moles of CaCO<sub>1</sub>
  - I confly
  - b. III only
  - c. IIII only
  - d. I and III only
  - e. I, II, and III

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