

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) =14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) +4(2) =14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) =6 O atoms

Product side

4 Fe atoms

2(3) =6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃:4 mol Fe

Stoichiometry Practice Problems Sheet Answer Key

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Table of Contents Stoichiometry Practice Problems Sheet Answer Key

1. Understanding the eBook Stoichiometry Practice Problems Sheet Answer Key
 - The Rise of Digital Reading Stoichiometry Practice Problems Sheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Practice Problems Sheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Practice Problems Sheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Practice Problems Sheet Answer Key
 - Personalized Recommendations
 - Stoichiometry Practice Problems Sheet Answer Key User Reviews and Ratings
 - Stoichiometry Practice Problems Sheet Answer Key and Bestseller Lists
5. Accessing Stoichiometry Practice Problems Sheet Answer Key Free and Paid eBooks
 - Stoichiometry Practice Problems Sheet Answer Key Public Domain eBooks
 - Stoichiometry Practice Problems Sheet Answer Key eBook Subscription Services
 - Stoichiometry Practice Problems Sheet Answer Key Budget-Friendly Options
6. Navigating Stoichiometry Practice Problems Sheet Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Stoichiometry Practice Problems Sheet Answer Key Compatibility with Devices
- Stoichiometry Practice Problems Sheet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Practice Problems Sheet Answer Key
 - Highlighting and Note-Taking Stoichiometry Practice Problems Sheet Answer Key
 - Interactive Elements Stoichiometry Practice Problems Sheet Answer Key
- 8. Staying Engaged with Stoichiometry Practice Problems Sheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Practice Problems Sheet Answer Key
- 9. Balancing eBooks and Physical Books Stoichiometry Practice Problems Sheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Practice Problems Sheet Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Practice Problems Sheet Answer Key
 - Setting Reading Goals Stoichiometry Practice Problems Sheet Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Practice Problems Sheet Answer Key
 - Fact-Checking eBook Content of Stoichiometry Practice Problems Sheet Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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