

Stoichiometry Worksheet #1

1. Aluminum chloride, AlCl_3 , is used as a catalyst in various industrial reactions. It is prepared from hydrogen chloride gas and aluminum metal shavings.



1A. How many moles of AlCl_3 can be prepared from 3.5 moles of hydrogen chloride gas with an excess of aluminum?

Ans: **1.2 mol AlCl_3**

1B. How many moles of hydrogen gas would be produced from the use of 8.5 moles of aluminum with an excess of hydrogen chloride?

Ans: **13 mol H_2**

2. When dinitrogen pentoxide, N_2O_5 , a white solid, is heated, it decomposes to nitrogen dioxide and oxygen.



2A. How many moles of nitrogen dioxide can be formed from the decomposition of 1.25 g of N_2O_5 ?

Ans: **0.0231 mol NO_2**

2B. How many grams of oxygen can be formed from the decomposition of 2.3 g of N_2O_5 ?

Ans: **0.34 g O_2**

3. Chlorine is prepared from sodium chloride by electrochemical decomposition. Formerly chlorine was produced by heating hydrochloric acid with pyrolusite [manganese dioxide or manganese(IV) oxide, MnO_2], a common manganese ore. Small amounts of chlorine may be prepared in the laboratory by the same reaction.



3A. How many grams of HCl react with 5.00 g of MnO_2 , according to the equation?

Ans: **8.39 g HCl**

3B. If a chemist wanted to prepare 100. g of chlorine, how many grams of MnO_2 are needed, assuming there is more than enough hydrochloric acid?

Ans: **123 g MnO_2**

3C. How many molecules of water are produced from the reaction of 5.0 g of HCl ?

Ans: **4.1×10^{23} molec. H_2O**

4. Sodium is a soft, reactive metal that instantly reacts with water to give hydrogen gas and a solution of sodium hydroxide, NaOH .

4A. How many grams of sodium metal are needed to give 7.81 g of hydrogen by this reaction? Ans: **178 g Na**

4B. How many sodium atoms are needed to react with 1.25×10^{24} molecules of water? Ans: **1.25×10^{24} atoms Na**

Stoichiometry Mixed Problems Worksheet 1

Daniel C. Tofan



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Table of Contents Stoichiometry Mixed Problems Worksheet 1

1. Understanding the eBook Stoichiometry Mixed Problems Worksheet 1
 - The Rise of Digital Reading Stoichiometry Mixed Problems Worksheet 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Mixed Problems Worksheet 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Mixed Problems Worksheet 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Mixed Problems Worksheet 1
 - Personalized Recommendations
 - Stoichiometry Mixed Problems Worksheet 1 User Reviews and Ratings
 - Stoichiometry Mixed Problems Worksheet 1 and Bestseller Lists
5. Accessing Stoichiometry Mixed Problems Worksheet 1 Free and Paid eBooks
 - Stoichiometry Mixed Problems Worksheet 1 Public Domain eBooks
 - Stoichiometry Mixed Problems Worksheet 1 eBook Subscription Services
 - Stoichiometry Mixed Problems Worksheet 1 Budget-Friendly Options

6. Navigating Stoichiometry Mixed Problems Worksheet 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Mixed Problems Worksheet 1 Compatibility with Devices
 - Stoichiometry Mixed Problems Worksheet 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Mixed Problems Worksheet 1
 - Highlighting and Note-Taking Stoichiometry Mixed Problems Worksheet 1
 - Interactive Elements Stoichiometry Mixed Problems Worksheet 1
8. Staying Engaged with Stoichiometry Mixed Problems Worksheet 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Mixed Problems Worksheet 1
9. Balancing eBooks and Physical Books Stoichiometry Mixed Problems Worksheet 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Mixed Problems Worksheet 1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Mixed Problems Worksheet 1
 - Setting Reading Goals Stoichiometry Mixed Problems Worksheet 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Mixed Problems Worksheet 1
 - Fact-Checking eBook Content of Stoichiometry Mixed Problems Worksheet 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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