Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)

- N₂ + 2O₂ → N₂O₄
 - a. If 15.0g of N₂O₄ was produced, how many moles of O₂ were required?

15.0g N₂O₄ 1 mol N₂O₄ 2 mol O₂ = 0.326 mol O₂ 92.0g N₂O₄ 1 mol N₂O₄

b. If 4.0x10⁻³ moles of oxygen reacted, how many grams of N₂ were needed?

4.0x10⁻³ mol O₂ 1 mol N₂ 28.0g N₂ = 5.6x10⁻² g N₂ 2 mol O₂ 1 mol N₂ = 5.6x10⁻² g N₂

- K₃PO₄ + AJ(NO₃)₃ → 3KNO₃ + AJPO₄
 - a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

2.04 mol K₃PO₄ 3 mol KNO₃ 101.1 g KNO₃ = 619g KNO₃ 1 mol K₃PO₄ 1 mol KNO₃

> b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

5.80g AIPO₄ 1 mol AIPO₆ 1 mol AI(NO₁)₁ = 4.75×10⁻² mol AI(NO₁)₁ 122.0g AIPO₆ 1 mol AIPO₆

- CaC₂ + 2H₂O → Ca(OH)₂ + C₂H₂
 - a. If you have 5.50 mol of CaC₂, how much C₂H₂ do you get?

5.50 mol CaC₂ 1 mol C₂H₂ 26.0g C₂H₂ = 143g C₂H₂ 1 mol CaC₂ 1 mol C₂H₂

b. How many moles of water are needed when 65.0g of CaC₂ have reacted?

65.0g CaC₂ 1 mol CaC₂ 2 mol H₂O = 2.03 mol H₂O 64.1g CaC₂ 1 mol CaC₂

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose (C₆H₁₂O₆). Write and balance the chemical equation. How many moles of CO₂ are required to make 120.0g of glucose?

6H2O + 6CO2 -- C4H22O4 + 6O2

120.0g C₆H₁₂O₈ 1 mol C₆H₁₂O₈ 6 mol CO₂ = 4.000 mol CO₂ 180.0 g C₆H₁₂O₈ 1 mol C₆H₁₂O₆

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