

Name: _____

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STOICHIOMETRY WS 10: Stoichiometry, Limiting Reactant and Percent Yield Problems

- 1) Carbon dioxide is a greenhouse gas that is linked to global warming. It is released into the atmosphere through the combustion of octane (C_8H_{18}) in gasoline. Write the balanced chemical equation for the combustion of octane and calculate the mass of octane needed to release 5.00 mol of CO_2 _____.

Balanced Equation: _____

- 2) When aluminum is mixed with iron (III) oxide, iron metal and aluminum oxide are produced, along with a large quantity of heat. What mole ratio would you use to determine moles of Fe if moles of Fe_2O_3 is known?



- 3) The vigorous reaction between potassium and water is shown below. How many moles of hydrogen are produced if only 0.0400 mol of potassium is used? _____



- 4) Methane and sulfur react to produce carbon disulfide (CS_2), a liquid often used in the production of cellophane. Balance the equation and then calculate the moles of CS_2 produced when 1.50 moles of S_8 are used. _____



- 5) Solid silver chromate (Ag_2CrO_4) forms when potassium chromate (the excess reactant in this problem) is added to a solution containing 0.500 g of silver nitrate. Determine the theoretical yield of silver chromate. Calculate the percent yield if the reaction produces 0.455 g of silver chromate. _____
(HINT: when you write the equation, it is a double replacement)

Balanced Equation: _____

Stoichiometry Ws F Mixed Problems Answers

Peipei Pang



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