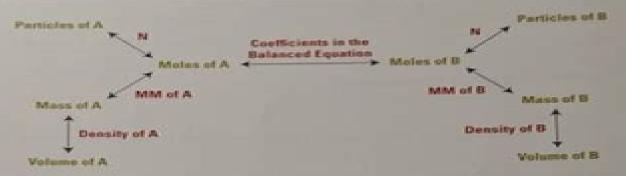
Stoichiometry

Name: Section:



1. Consider the following balanced reaction equation:

- a. What are the following molar ratios:
 - i. mol Na₂CO₃/mol H₃PO₃
 - ii. mol Na₂CO₂/mol H₂CO₃
 - iii. mol Na₃PO₃/mol H₂CO₃ _____
 - iv. mol H₃PO₃/mol Na₃PO₃
- Calculate the number of moles of carbonic acid that can be made from 0.833 moles phosphorous acid.
- c. How many moles of sodium carbonate are needed to react with 3.55 × 10⁻³ mol of phosphorous acid?
- d. How many moles sodium phosphite can be produced with 0.075 moles of carbonic acid?

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