

Stoichiometry



Name: _____

Section: _____



1. Consider the following balanced reaction equation:



- a. What are the following molar ratios:
 - i. mol Na_2CO_3 /mol H_3PO_3 _____
 - ii. mol Na_2CO_3 /mol H_2CO_3 _____
 - iii. mol Na_3PO_3 /mol H_2CO_3 _____
 - iv. mol H_3PO_3 /mol Na_3PO_3 _____
- b. Calculate the number of moles of carbonic acid that can be made from 0.833 moles phosphorous acid.
- c. How many moles of sodium carbonate are needed to react with 3.55×10^{-3} mol of phosphorous acid?
- d. How many moles sodium phosphite can be produced with 0.075 moles of carbonic acid?

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Stoichiometry Unit Project Luann Marie Decker, 1998

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