

Stoichiometry Worksheet #1

1. Aluminum chloride, AlCl_3 , is used as a catalyst in various industrial reactions. It is prepared from hydrogen chloride gas and aluminum metal shavings.



1A. How many moles of AlCl_3 can be prepared from 3.5 moles of hydrogen chloride gas with an excess of aluminum?

$$3.5 \text{ mol HCl} \times \frac{2 \text{ mol AlCl}_3}{6 \text{ mol HCl}} = 1.2 \text{ mol AlCl}_3$$

Ans: 1.2 mol AlCl_3

1B. How many moles of hydrogen gas would be produced from the use of 8.5 moles of aluminum with an excess of hydrogen chloride?

Ans: 13 mol H_2

$$8.5 \text{ mol Al} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} = 13 \text{ mol H}_2$$

2. When dinitrogen pentoxide, N_2O_5 , a white solid, is heated, it decomposes to nitrogen dioxide and oxygen.



2A. How many moles of nitrogen dioxide can be formed from the decomposition of 1.25 g of N_2O_5 ?

$$1.25 \text{ g N}_2\text{O}_5 \times \frac{1 \text{ mol N}_2\text{O}_5}{108.02 \text{ g N}_2\text{O}_5} \times \frac{4 \text{ mol NO}_2}{2 \text{ mol N}_2\text{O}_5} = 0.0231 \text{ mol NO}_2$$

Ans: 0.0231 mol NO_2

2B. How many grams of oxygen can be formed from the decomposition of 2.3 g of N_2O_5 ?

$$2.3 \text{ g N}_2\text{O}_5 \times \frac{1 \text{ mol N}_2\text{O}_5}{108.02 \text{ g N}_2\text{O}_5} \times \frac{1 \text{ mol O}_2}{2 \text{ mol N}_2\text{O}_5} \times \frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} = 0.34 \text{ g O}_2$$

Ans: 0.34 g O_2

3. Chlorine is prepared from sodium chloride by electrochemical decomposition. Formerly chlorine was produced by heating hydrochloric acid with pyrolusite [manganese dioxide or manganese(IV) oxide, MnO_2], a common manganese ore. Small amounts of chlorine may be prepared in the laboratory by the same reaction.



3A. How many grams of HCl react with 5.00 g of MnO_2 , according to the equation?

$$5.00 \text{ g MnO}_2 \times \frac{1 \text{ mol MnO}_2}{86.94 \text{ g MnO}_2} \times \frac{4 \text{ mol HCl}}{1 \text{ mol MnO}_2} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 8.39 \text{ g HCl}$$

Ans: 8.39 g HCl

3B. If a chemist wanted to prepare 100. g of chlorine, how many grams of MnO_2 are needed, assuming there is more than enough hydrochloric acid?

Ans: 123 g MnO_2

$$100. \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.90 \text{ g Cl}_2} \times \frac{1 \text{ mol MnO}_2}{1 \text{ mol Cl}_2} \times \frac{86.94 \text{ g MnO}_2}{1 \text{ mol MnO}_2} = 123 \text{ g MnO}_2$$

3C. How many molecules of water are produced from the reaction of 5.0 g of HCl?

Ans: 4.1×10^{23} molec. H_2O

$$5.0 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.46 \text{ g HCl}} \times \frac{2 \text{ mol H}_2\text{O}}{4 \text{ mol HCl}} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 4.1 \times 10^{23} \text{ molecules H}_2\text{O}$$

4. Sodium is a soft, reactive metal that instantly reacts with water to give hydrogen gas and a solution of sodium hydroxide,



4A. How many grams of sodium metal are needed to give 7.81 g of hydrogen by this reaction?

Ans: 178 g Na

$$7.81 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2.02 \text{ g H}_2} \times \frac{2 \text{ mol Na}}{1 \text{ mol H}_2} \times \frac{22.99 \text{ g Na}}{1 \text{ mol Na}} = 178 \text{ g Na}$$

4B. How many sodium atoms are needed to react with 1.25×10^{24} molecules of water?

Ans: 1.25×10^{24} atoms Na

$$1.25 \times 10^{24} \text{ molecules H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{6.02 \times 10^{23} \text{ molecules H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ atoms Na}}{1 \text{ mol Na}} =$$

Stoichiometry Problem Solving Workbook Answers

Pamela Mills, Amina El-Ashmawy



Stoichiometry Problem Solving Workbook Answers:

Problems and Problem Solving in Chemistry Education Georgios Tsaparlis, 2021 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner *General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material *Class 8-12 Chemistry Questions and Answers PDF* Arshad Iqbal, The Class 8 12 Chemistry Quiz Questions and Answers PDF Grade 8 12 Chemistry Competitive Exam Questions Chapter 1 15 Practice Tests Chemistry Textbook Questions for Beginners includes Questions to solve problems with hundreds of class questions Class 8 12 Chemistry Questions and Answers PDF book covers basic concepts and analytical assessment tests Class 8 12 Chemistry Quiz PDF book helps to practice test questions from exam prep notes The Grade 8 12 Chemistry Quiz Questions and Answers PDF eBook includes Practice material with verbal quantitative and analytical past papers questions Class 8 12 Chemistry Questions and Answers PDF Free download chapter 1 a book to review textbook questions on chapters Molecular structure acids and bases atomic structure bonding chemical equations descriptive chemistry equilibrium systems gases laboratory liquids and solids mole concept oxidation reduction rates of reactions solutions thermochemistry Questions for high school and college revision questions Chemistry Interview Questions and Answers PDF Download free eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The Grade 8 12 Chemistry Interview Questions Chapter 1 15 PDF book includes high school workbook questions to

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