- Remember, this is practice Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice
  without a calculator, even if it slows you down. Look for the "easy math" common factors and rough estimation do
  not do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.
- 1. Balance the following equation:

 $NH_1 + O_2 \rightarrow NO_2 + H_2O$ 

The balanced equation shows that 1.00 mole of NHs requires \_\_\_ mole(s) of O<sub>2</sub>

- 0.57
- b. 1.25
- c. 1.33
- d. 1.75
- e. 3.5
- Write a balanced equation for the combustion of acetaldehyde, CH<sub>2</sub>CHO.

When properly balanced, the equation indicates that mole(s) of O<sub>2</sub> are required for each mole of CH<sub>2</sub>CHO.

- m. 1
- bc 2:
- c. 2.5
- d. 3
- e. 5
- What is the total mass of products formed when 16 grams of CH<sub>4</sub> is burned with excess oxygen?
  - in. 32 m.
  - b. 36 gr
  - c. 44 g.
  - d. 62 g
  - c. 80 g
- Write a balanced equation for the combustion of propane, C<sub>3</sub>H<sub>8</sub>.

When balanced, the equation indicates that \_\_\_\_ moles of O<sub>2</sub> are required for each mole of C<sub>3</sub>H<sub>6</sub>.

- m. 1.5
- ъ. з.
- c. 3.5
- dl. 5
- er. 80

 Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the <u>sum</u> of the coefficients in the balanced equation:

 $KClO_3 \rightarrow KCl + O_2$ 

- a. 3
- b. 5
- c. 6
- d. 7
- e. 8
- Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below.

- a. 1.5 g
- b. 2.0 g
- c. 3.0 g
- d. 6.0 g.
- e. 12 gr
- How many grams of nitric acid, HNO<sub>3</sub>, can be prepared from the reaction of 138 g of NO<sub>2</sub> with 54.0 g H<sub>2</sub>O according to the equation below?

$$3NO_2 + H_2O \rightarrow 2HNO_1 + NO$$

- m. 92
- 6 108.
- c. 126
- d. 189
- e. 279
- 8. Which of the following statements is true?
  - The molar mass of CaCO<sub>3</sub> is 100.1 g mol<sup>-1</sup>.
  - 50 g of CaCO<sub>3</sub> contains 9 × 10<sup>23</sup> oxygen atoms.
  - III. A 200 g sample of CaCO<sub>1</sub> contains 2 moles of CaCO<sub>1</sub>
  - I confly
  - b. III only
  - c. IIII only
  - d. I and III only
  - e. I, II, and III

# **Stoichiometry Practice Test Answers**

**Bill T Reese** 

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web 33 jacob looked up and there was esau coming with his four hundred men so he divided the children among leah rachel and the two female servants 2 he put the female servants and their children in front leah and her children next and rachel and joseph in the rear 3 he himself went on ahead and bowed down to the ground seven times as he appr

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web according to the hebrew bible esau is the progenitor of the edomites and the elder brother of jacob the patriarch of the israelites 8 jacob and esau were the sons of isaac and rebecca and the grandsons of abraham and sarah of the twins esau was the first to be born with jacob following holding his heel

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web the story focuses on esau s loss of his birthright to jacob and the conflict that ensued between their descendant nations because of jacob s deception of their aged and blind father isaac in order to receive esau s birthright blessing from isaac jacob and esau the church of jesus christ of latter day saints - Nov 06 2022

web the lord wanted jacob to have the birthright blessing because esau was not worthy of it jacob asked esau to trade him the birthright for some food esau agreed and traded his birthright to jacob genesis 25 23 29 34 hebrews 11 20 image rebekah and isaac rebekah and isaac wanted what was best for their children

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