

Limiting Reactants & Percent Yield – Ch. 9

1. Identify the limiting and excess reagents when 25 g of nitrogen reacts with 25 g of hydrogen. How many grams of ammonia gas are formed in this reaction?



2. 50.0 g H_2SO_4 react with 75.0 g NaOH . Identify the limiting and excess reactants. How many grams of Na_2SO_4 will be formed?



3. If 6.57 g of iron are reacted with an excess of hydrochloric acid, HCl , then hydrogen gas and 14.63 g of iron(III) chloride are obtained. Calculate the theoretical yield and percent yield of FeCl_3 .



4. A chemist burns 160. g of aluminum in oxygen to produce aluminum oxide. She produces 260. g of aluminum oxide. Calculate the theoretical yield and percent yield.



Stoichiometry Limiting Reactant Percent Yield Quiz

JL Elias



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